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Passage 1 (Questions 1-4)

To test how amino acids first appeared on this planet, an experiment was set up to simulate the conditions thought to have existed on early Earth. H_2O , CH_4 , NH_3 , and H_2 were all introduced into a sterile arrangement of sealed flasks and tubes. Liquid water evaporated in one compartment as "lightning" was induced between a pair of electrodes. The water vapor then cooled and condensed in a connected section, which fed back to the first compartment, until a dynamic equilibrium was reached.

After two weeks of continuous operation, roughly 2% of the carbon in the system was in the form of amino acids, with glycine being the most abundant. The following steps are believed to account for the formation of the key intermediates, formaldehyde and hydrogen cyanide.



Mechanism 1

Via the Strecker Synthesis, glycine can then form as follows:



Mechanism 2

Several other amino acids were also observed to form. Of the chiral amino acids, both enantiomers were created in a racemic mixture, though in most biological systems L-amino acids are used exclusively for protein synthesis.

Simple sugars also formed in the same environment. Once formaldehyde is present to react with H_2O , it does so according to the Butlerov reaction:



Figure 1. Formation of carbohydrates from aqueous formaldehyde

Question 4

Which of the following will increase the yield of carboxylic acid formed from a nitrile group-containing compound, using a mechanism similar to the Strecker synthesis?

- ☐ A. Use a stronger nucleophile, like a Grignard Reagent, instead of H_2O
- ☐ B. Attach an electron withdrawing group to the carbon of the nitrile group
- ☐ C. Use a stronger electrophile, like a Grignard Reagent, instead of H_2O
- ☐ D. Attach an electron donating group to the carbon of the nitrile group

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Annelies Wilder-Smith



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