

SIGNIFICANT FIGURE WORKSHEET-ANSWER KEY

1) Indicate the number of significant figures in the following:

- | | | | | | |
|---------------|----------|----------------|----------|--------------------------|----------|
| a) 2.508 L | <u>4</u> | d) 12 in/1 ft | <u>∞</u> | g) 29 mm | <u>2</u> |
| b) 0.00210 km | <u>3</u> | e) 20 students | <u>∞</u> | h) 3.50×10^3 nm | <u>3</u> |
| c) 3000 g | <u>1</u> | f) 45.00 g | <u>4</u> | i) 8.34:5 g/mL | <u>3</u> |

2) Round each of the following numbers to the indicated number of significant figures:

- | | | | |
|-----------------------------------|----------------------------|---------------------------------------|--|
| a) 96.528 cm ² to 3 SF | <u>96.5 cm²</u> | d) 0.036559 m to 4 SF | <u>0.03656 m</u> |
| b) 1275 mL to 2 SF | <u>1300 mL</u> | e) 0.14500001 lb to 2 SF | <u>0.15 lb</u> |
| c) 26.485 g to 4 SF | <u>26.48 g</u> | f) 6.0221367×10^{23} to 4 SF | <u>6.022×10^{23}</u> |

3) Perform the following calculations, giving the answer properly rounded to the correct number of significant figures.

- | | | | |
|---|----------------|---|---------------|
| a) $65\text{g} + 25.86\text{g} + 145.1\text{g} =$ | <u>236 g</u> | c) $2.5 \times 10^2 + 2.6 \times 10^3 + 3.4 \times 10^{-1} =$ | <u>2900</u> |
| b) $415.1\text{g} - 25.56\text{g} =$ | <u>389.5 g</u> | d) $\frac{4.582\text{g} + 4.54\text{g} + 4.60\text{g}}{3} =$ | <u>4.57 g</u> |

4) What is the assumed absolute uncertainty (au) of the following?

- | | | | |
|--------------------------|---|--------------------------------|---------------------------------------|
| a) 0.0770 g | <u>$\pm 0.0001\text{ g}$</u> | d) $1.56 \times 10^3\text{ m}$ | <u>$\pm 10\text{ m}$</u> |
| b) 450 mL | <u>$\pm 10\text{ mL}$</u> | e) 4500 km | <u>$\pm 100\text{ km}$</u> |
| c) 35:14 cm ² | <u>$\pm 100\text{ cm}^2$</u> | f) 25.0 lb | <u>$\pm 0.1\text{ lb}$</u> |

5) What is the assumed relative uncertainty (ru) of the following?

- | | | | |
|--------------------------|--------------|--------------------------------|--------------|
| a) 0.0770 g | <u>0.001</u> | d) $1.56 \times 10^3\text{ m}$ | <u>0.006</u> |
| b) 450 mL | <u>0.02</u> | e) 4500 km | <u>0.02</u> |
| c) 35:14 cm ² | <u>0.03</u> | f) 25.0 lb | <u>0.004</u> |

6) Perform the following calculations, giving the answer properly rounded to the correct number of significant figures.

- | | | | |
|--|----------------|--|--|
| a) $6.72\text{ m} \times \frac{1\text{ mm}}{10^{-3}\text{ m}} =$ | <u>6720 mm</u> | c) $\frac{(8.015 \times 10^{-3})(0.550)}{1.9 \times 10^3} =$ | <u>2.3×10^{-6}</u> |
| b) $\frac{60.00 \times 35}{0.0050} =$ | <u>420,000</u> | d) $(22.5\text{ cm})(125.6\text{ cm})(1.4\text{ cm}) =$ | <u>40:00 cm³</u> |

7) Perform the following calculations, giving the answer properly rounded to the correct number of significant figures. Hint: You might want to show an intermediate answer.

- a) $\frac{35.46\text{ g/mol} - 35.4527\text{ g/mol}}{35.4527\text{ g/mol}} \times 100\% =$ 0.02:1 %

Mastering Chemistry Answer Key Significant Figures

Katrin Zwirgmaier



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