23. When this reaction comes to equilibrium, will the concentrations of the reactants or products be greater? Does the answer to this question depend on the initial concentrations of the reactants and products?

reactants will be in higher concentration b/c k is much
$$A(g) + B(g) = 2C(g)$$
 $K = 1.4 \times 10^{-5}$ Smaller than 1.

27. This reaction has an equilibrium constant of $K_p = 2.26 \times 10^4$ at 298 K:

$$CO(g) + 2 H_2(g) \rightleftharpoons CH_3OH(g)$$

Calculate K_p for each reaction and predict whether reactants or products will be favored at equilibrium:

a.
$$CH_3OH(g) = CO(g) + 2H_2(g) / 2.26e4 - 4.42e-5$$

b.
$$\frac{1}{2}CO(g) + H_2(g) = \frac{1}{2}CH_3OH(g) (2.26e4)^{4} \cdot 150.33$$

c.
$$2 \text{ CH}_3 \text{OH}(g) = 2 \text{ CO}(g) + 4 \text{ H}_2(g) / (2.26e4)^{1/2} = 6.65e-3$$

Write an equilibrium expression for each chemical equation.

a.
$$CO_3^{2-}(aq) + H_2O(l) = HCO_3^{-}(aq) + OH^{-}(aq)$$

c.
$$HF(aq) + H_2O(l) \Rightarrow H_3O^+(aq) + F^-(aq) \times G^-3Ch_3O^-3$$

d.
$$NH_3(aq) + H_2O(l) = NH_4^+(aq) + OH^-(aq) V \cdot \frac{DH_4^+3(OH)}{V \cdot DH_4^+3(OH)}$$

35. Consider the reaction:

$$CO(g) + 2 H_2(g) = CH_2OH(g)$$

$$V = \frac{CCH_2OH7}{CCOT(H_2)^2} = \frac{(0.185)}{(0.105)(0.00)^2} = 136$$

An equilibrium mixture of this reaction at a certain temperature has $[CO] = 0.105 M_{\bullet}[H_2] = 0.114 M_{\bullet}$ and $[CH_1OH] = 0.185 M_{\bullet}$ What is the value of the equilibrium constant (K_c) at this temperature?

Mastering Chemistry Exercise Answers Ch 15

Tom Cole

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put simply the four quadrants that the controller can work in are 1 motor going clockwise controller applying clockwise drive 2 motor going anti clockwise controller applying anti clockwise drive 3 motor going clockwise controller applying anti clockwise drive 4 motor going anti clockwise controller applying clockwise drive

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