

**ENGAGE**

Why does the atomic number but not the mass number change in positron emission?

**SAMPLE PROBLEM 5.4** Writing a Nuclear Equation for Positron Emission**TRY IT FIRST**

Write the balanced nuclear equation for manganese-49, which decays by emitting a positron.

**SOLUTION GUIDE**

	Given	Need	Connect
<b>ANALYZE THE PROBLEM</b>	Mn-49, positron emission	balanced nuclear equation	mass number, atomic number of new nucleus

**STEP 1** Write the incomplete nuclear equation.



**STEP 2** Determine the missing mass number. In the equation, the mass number, 49, is equal to the sum of the mass numbers of the new nucleus and the positron.

$$49 = ? + 0$$

$$49 - 0 = ?$$

$$49 - 0 = 49 \text{ (mass number of new nucleus)}$$

**STEP 3** Determine the missing atomic number. The atomic number, 25, must equal the sum of the atomic numbers of the new nucleus and the positron.

$$25 = ? + 1$$

$$25 - 1 = ?$$

$$25 - 1 = 24 \text{ (atomic number of new nucleus)}$$

**STEP 4** Determine the symbol of the new nucleus. On the periodic table, the element that has atomic number 24 is chromium, Cr. The atomic symbol for this isotope of Cr is written  ${}_{24}^{49}\text{Cr}$ .

**STEP 5** Complete the nuclear equation.

**STUDY CHECK 5.4**

Write the balanced nuclear equation for xenon-118, which undergoes positron emission.

**ANSWER****TEST**

Try Practice Problems 5.15 to 5.18

# Mastering Chemistry Exercise Answers

**CHANG**



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