

## Chemistry First Semester Review

## Introduction to Chemistry

## Terms to Define

- qualitative data
- quantitative data
- independent variable
- dependent variable
- control
- constants
- scientific theory
- scientific law
- standard of measurement
- base unit
- derived unit
- accuracy
- precision

Identify the following examples of data as qualitative (QL) or quantitative (QT).

- QL 1. the solution is blue  
 QL 2. the metal has luster  
 QN 3. the mass is 105 g  
 QN 4. the temperature is 12 °C  
 QN 5. the length is 5.0 mm  
 QL 6. the liquid conducts electricity

Identify the SI base unit used for each of the following:

time	s
length	m
mass	kg
temperature	K
amount of a substance	mol

## Metric Conversions

convert	to	
0.0032 mg	g	$3.2 \times 10^{-6}$
2390 km	m	$2.39 \times 10^6$
0.009 g	ng	$9 \times 10^6$
0.033 Mg	cg	$3.3 \times 10^6$
0.00032 s	ms	0.32
3.42 µg	pg	$3.42 \times 10^6$

## Temperature Conversions

Convert the following:

- 1)  $362\text{ °C} = 635\text{ K}$   
 2)  $25\text{ °C} = 298\text{ K}$   
 3)  $300\text{ K} = 27\text{ °C}$   
 4)  $0\text{ K} = -273\text{ °C}$   
 5)  $25\text{ K} = -248\text{ °C}$

## Experimental Design

You design an experiment to determine which household cleaner best prevents the growth of bacteria. You gather five different household cleaners and some petri dishes with bacteria samples. Determine each of the following for your experiment:

independent variable: type of cleaner  
 dependent variable: how well cleaners control bacteria  
 control: petri dish w/ bacteria, no cleaner  
 constants: temp, strain of bacteria, sunlight, etc

## Scientific Notation

Convert the following to scientific or standard notation:

- 1)  $8,320,000 = 8.32 \times 10^6$   
 2)  $0.000937 = 9.37 \times 10^{-4}$   
 3)  $7,500,000,000 = 7.5 \times 10^9$   
 4)  $5.22 \times 10^4 = 52,200$   
 5)  $9.214 \times 10^6 = 9,214,000$   
 6)  $8.5 \times 10^{-8} = 0.00000085$

## Scientific Notation

Perform the following calculations:

- 1)  $(2.1 \times 10^5) \times (8.33 \times 10^4) = 1.7 \times 10^{10}$   
 2)  $(7.50 \times 10^{30}) / (4.212 \times 10^{30}) = 1.78 \times 10^{-3}$   
 3)  $(4.4 \times 10^6) + (7.2 \times 10^6) = 5.1 \times 10^6$   
 4)  $(5.29 \times 10^9) - (6.04 \times 10^9) = -5.51 \times 10^9$

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magnetic quantum number mass of electron mass spectrometer metallic crystals properties metallic solids metals structure molar volume molecular ions molecular solids molecules moles Moseley law neutron properties non ideal behavior of gases orbital concept partial pressure calculations phase changes energies photons wave number Planck s quantum theory plasma state positive and negative ions pressure units properties of cathode rays covalent crystals properties of crystalline solids properties of positive rays quantum numbers quantum theory relative abundance Rutherford model of atom shapes of orbitals solid iodine structure solids properties solvent extraction spectrometer spin quantum number states of matter stoichiometry sublimation thermometry scales types of solids unit cell van der Waals equation vapor pressure and spectrum

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