

CHEMICAL THERMODYNAMICS

- 1) In open system, system and surroundings exchange.....
 (a) Energy only (b) Matter only
(c) Energy and matter (d) Neither energy nor matter
- 2) An isothermal expansion takes place at.....
 (a) constant pressure (b) constant volume
(c) constant temperature (d) constant temperature & pressure
- 3) In a closed thermos flask few cubes of ice are placed. It is example of.....
 (a) closed system (b) open system
(c) isolated system (d) thermodynamics system
- 4) In an adiabatic compression, the internal energy of the system.....
 (a) decreases **(b) increases** (c) increases or decreases (d) none
- 5) Identify the intensive quantities from the following.....
 (a) volume (b) enthalpy **(c) pressure** (d) none
- 6) Which of the following is extensive property?
 (a) mole (b) mass (c) volume **(d) all are correct**
- 7) The heat change in a chemical reaction at constant pressure is.....
(a) ΔH (b) ΔU (c) ΔT (d) ΔV
- 8) ΔU is always positive when.....
(a) system absorbs heat and work is done on it
 (b) system emits heat and work is done on it
 (c) system emits heat and no work is done on it
 (d) system absorbs heat and work is done by it
- 9) If in a chemical reaction, the products have less energy than the reactants, that reaction is.....
 (a) a reversible reaction **(b) an exothermic reaction**
 (c) an isothermal reaction (d) an endothermic reaction
- 10) A small piece of Na metal is dropped into water. Na bursts into flame, the reaction is:
 (a) neutralisation (b) reversible **(c) exothermic** (d) endothermic
- 11) For the reaction: $C_3H_8(g) + 5O_2(g) \rightarrow 3CO_2(g) + 4H_2O(l)$ at constant temperature, $\Delta H - \Delta U$ is
 (a) $+RT$ **(b) $-3RT$** (c) $+3RT$ (d) $-RT$
- 12) When NH_4Cl is dissolved in water, the solution becomes cold. The change is?
(a) endothermic (b) exothermic (c) super cooling (d) none

Mcqs Of Thermodynamics With Answers

Michael Brown



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Comprehensive Objective Physics Narinder Kumar, Dr. J. K. Juneja, 2006

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Problems In Chemical Thermodynamics, With Solutions Maka Aleksishvili, Shota Sidamonidze, 2002-11-26 The methods of chemical thermodynamics are effectively used in many fields of science and technology Mastering these methods and their use in practice requires profound comprehension of the theoretical questions and acquisition of certain calculating skills This book is useful to undergraduate and graduate students in chemistry as well as chemical thermal and refrigerating technology it will also benefit specialists in all other fields who are interested in using these powerful methods in their practical activities

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model question papers at the end of the book **Thermodynamics** Rumi Michael Leigh, 2023-05-24 Thermodynamics Things You Should Know Questions and Answers explains the principles of thermodynamics in a question and answer format The book introduces energy work heat and the basic laws of thermodynamics It describes the concepts of entropy enthalpy internal energy and equilibrium and shows how these ideas connect to physical and chemical processes Key topics include the first and second laws of thermodynamics the role of entropy in natural systems and the behavior of gases as described by equations of state The text also examines heat transfer efficiency and applications of thermodynamics in engines refrigeration and power systems Worked examples and explanations show how thermodynamic concepts are applied in solving problems and understanding physical phenomena This book will interest students engineering learners and those studying physics who want to understand thermodynamics and its applications Problems And Solutions On Thermodynamics And Statistical Mechanics Yung-kuo Lim, 1990-02-01 The material for these volumes has been selected from the past twenty years examination questions for graduate students at University of California at Berkeley Columbia University the University of Chicago MIT State University of New York at Buffalo Princeton University and University of Wisconsin

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y hasta el punto de que afirmar que cumbres borrascosas es una novela de amor resulta un pobre esquematismo sólo con el juego de narradores que se trae este relato bastaría para indicar la complejidad de sus pretensiones el principal es un narrador identificado en la figura de un forastero urbano que se deja caer por los páramos con la

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