

Iodine Clock Reaction Lab - Data Tables

1. Find the Volume of a Drop of Solution

Mass of first five drops of water	0.03g
Average mass of one drop of water	.006g
Mass of second five drops of water	0.04g
Average mass of one drop of water	.008g
Mass of third five drops of water	0.08g
Average mass of one drop of water	.008g
Average of masses of one drop of water	.0053
Volume of one drop of water	5.3×10^{-6}

2. Determine the Reaction Rate and Calculate the Rate Law

	Time (s)			Temp (°C)	Reaction Rate (M/s)	Initial Concentrations, M		
	Trial 1	Trial 2	Average			[I ⁻]	[BrO ₃ ⁻]	[H ⁺]
Exp. 1	12.97	14.30	16.18	20.0	8.9×10^{-5}	.0017	.0067	.0167
Exp. 2	6.50	7.05	6.78	20.0	2.05×10^{-4}	.0033	.0067	.0167
Exp. 3	4.29	3.74	4.02	20.0	3.46×10^{-5}	.0017	.00200	.0167
Exp. 4	4.59	5.11	4.85	20.0	2.86×10^{-5}	.0017	.0067	.0333
Exp. 5	4.91	2.05	3.48	20.0	3.44×10^{-5}	.0028	.0100	.0250

Sample Calculations:

$$\text{Rate} = \frac{5.3 \times 10^{-6} \text{ mol/L} \cdot \text{s}}{6} = \frac{8.9 \times 10^{-5}}{(2.5 \times 10^{-6})} = 1.39 \times 10^{-5} \text{ M/s}$$

Rate of Reaction	$\frac{1.39 \times 10^{-5}}{16.2} =$
Initial Concentration	$.0167 \cdot 2 = \frac{.02}{2} = .0017 \text{ M I}^-$

Iodine Clock Experiment Lab Report

Brendan G. Carr



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