

## Ch 13 Study Guide (Chemical Bonds)

1. Fill in the table below with the correct values:

Element Name	Ion Symbol	# Protons	# Electrons	Charge	Ion Type
Fluorine	$F^-$	9	10	-1	Anion
Iodine	$I^-$	53	54	-1	Anion
Sulfur	$S^{2-}$	16	18	-2	Anion
Potassium	$K^+$	19	18	+1	Cation
Calcium	$Ca^{2+}$	20	18	+2	Cation
Bromine	$Br^-$	35	36	-1	Anion
Strontium	$Sr^{2+}$	38	36	+2	Cation
Oxygen	$O^{2-}$	8	10	-2	Anion
Magnesium	$Mg^{2+}$	12	10	+2	Cation
Aluminum	$Al^{3+}$	13	10	+3	Cation
Selenium	$Se^{2-}$	34	36	-2	Anion
Lithium	$Li^+$	3	2	+1	Cation
Rubidium	$Rb^+$	37	36	+1	Cation
Chlorine	$Cl^-$	17	18	-1	Anion

2. An ionic bond is formed between what two types of atoms?

Metal & nonmetal

3. A covalent bond is formed between what two types of atoms?

Nonmetal & Nonmetal

4. Draw the bond(s) that will form between Mg and O.



5. What is the formula for the compound formed in question #4?



# Ion Activity Packet Answers

**McGraw-Hill Staff**



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