# IONIC BOND

# **Definition & Key Concepts**



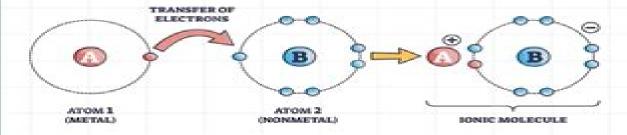
#### Definition

An ionic bond is a chemical bond formed due to the transfer of electrons from one atom to another, resulting in oppositely charged ions that attract each other.

 Example: NaCl (Sodium Chloride) Table Salt

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lonic bonding concepts evolved from Faraday's electrolysis experiments, later refined by Lewis and Kossel, who proposed atoms transfer electrons to achieve stability.



#### HOW DOES IT FORM?

- Electron Transfer: One atom donates an electron, another gains it.
- 2 Ion Formation: The donor becomes a positively charged cation (e.g., Na\*), and the receiver becomes a negatively charged anion (e.g., Cl-).
- S Electrostatic Attraction: Oppositely charged ions attract, forming a strong bond.

#### CHARACTERISTICS OF IONIC BOND

- Electron Transfer: Complete transfer of electrons.
- Electrostatic Attraction: Strong attraction between oppositely charged ions.
- Ion Formation: Forms cations (+) and anions (-).
- Crystal Lattice: Arranges in a rigid lattice structure.
- High Melting & Boiling Points: Due to strong ionic bonds.
- Conductivity: Conducts electricity in molten or dissolved state, not in solid.

#### COMMON EXAMPLES OF IONIC COMPOUNDS

- Sodium Chloride (NaCl) Table Salt
- Potassium Bromide (KBr) Used in Photography
- Nagnesium Oxide (MgO) Found in Medicines

# **Ionic Bond Review And Reinforce**

Michael J. Padilla

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