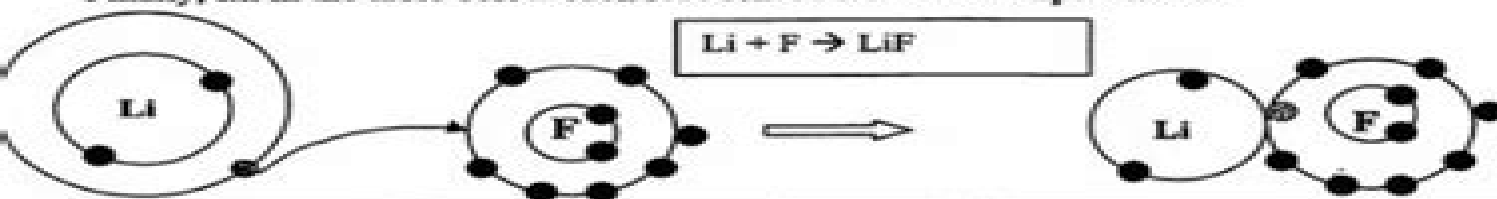


Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
Li	1	1 to F	Li^+
F	7	1 from Li	F^-

Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) $\text{Li} + \text{Cl} \Rightarrow \text{LiCl}$	Li	1	1 to Cl	Li^+
	Cl	7	1 from Li	Cl^-
2) $\text{Ca} + \text{O} \Rightarrow \text{CaO}$	Ca	2	2 to O	Ca^{2+}
	O	6	2 from Ca	O^{2-}
3) $\text{Be} + \text{F} \Rightarrow \text{BeF}_2$	Be	2	1 to each F	Be^{2+}
	F	7	Each 1 gets 1 from Be	F^-
4) $\text{Mg} + \text{S} \Rightarrow \text{MgS}$	Mg	2	2 to S	Mg^{2+}
	S	6	2 from Mg	S^{2-}
5) $\text{K} + \text{F} \Rightarrow \text{KF}$	K	1	1 to F	K^+
	F	7	1 from K	F^-

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Table of Contents Ionic Covalent Bonding Worksheets With Answers

1. Understanding the eBook Ionic Covalent Bonding Worksheets With Answers
 - The Rise of Digital Reading Ionic Covalent Bonding Worksheets With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Covalent Bonding Worksheets With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ionic Covalent Bonding Worksheets With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic Covalent Bonding Worksheets With Answers
 - Personalized Recommendations
 - Ionic Covalent Bonding Worksheets With Answers User Reviews and Ratings
 - Ionic Covalent Bonding Worksheets With Answers and Bestseller Lists

5. Accessing Ionic Covalent Bonding Worksheets With Answers Free and Paid eBooks
 - Ionic Covalent Bonding Worksheets With Answers Public Domain eBooks
 - Ionic Covalent Bonding Worksheets With Answers eBook Subscription Services
 - Ionic Covalent Bonding Worksheets With Answers Budget-Friendly Options
6. Navigating Ionic Covalent Bonding Worksheets With Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Ionic Covalent Bonding Worksheets With Answers Compatibility with Devices
 - Ionic Covalent Bonding Worksheets With Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ionic Covalent Bonding Worksheets With Answers
 - Highlighting and Note-Taking Ionic Covalent Bonding Worksheets With Answers
 - Interactive Elements Ionic Covalent Bonding Worksheets With Answers
8. Staying Engaged with Ionic Covalent Bonding Worksheets With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ionic Covalent Bonding Worksheets With Answers
9. Balancing eBooks and Physical Books Ionic Covalent Bonding Worksheets With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ionic Covalent Bonding Worksheets With Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ionic Covalent Bonding Worksheets With Answers
 - Setting Reading Goals Ionic Covalent Bonding Worksheets With Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ionic Covalent Bonding Worksheets With Answers
 - Fact-Checking eBook Content of Ionic Covalent Bonding Worksheets With Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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