

How to Name Chemical Compounds

is it ionic (usually),
or does it contain a polyatomic ion (as provided by H⁺)?

Ionic compound

[cation's name] [anion's name]

main- element	<ul style="list-style-type: none"> take its element name! ex. Al^{3+}: aluminum if transition metal*, include charge in Roman numeral ex. Cr^{2+}: chromium (II) 	<ul style="list-style-type: none"> get element's name, chop end, add "-ide" ex. F^{-}: fluorine → fluoride ✓
poly- element	<ul style="list-style-type: none"> just use its name! ex. NH_4^{+}: ammonium 	<ul style="list-style-type: none"> just use its name! ex. CO_3^{2-}: carbonate

* exceptions (no Roman numerals):

silver ion: always Ag^{+}

zinc ion: always Zn^{2+}

cadmium ion: always Cd^{2+}

Molecular compound

is it HClO_4 ?

Acid! then use nomenclature

[#]-[1st element name] [#]-[2nd element root]-"-ide"

↑ skip if "mono-" ↑ skip if 1st element is H

Numerical prefixes

- 1: mono-
- 2: di-
- 3: tri-
- 4: tetra-
- 5: penta-
- 6: hexa-
- 7: hepta-
- 8: octa-
- 9: nona-
- 10: deca-

for
first element: never
ex. N_2O_5 → dinitrogen
pentoxide

Acids

Ignore all H^{+} , and look at the name of the remaining anion.

Anion	X-ide	X-ate	X-ite
Acid	hydro-X-ic acid	X-ic acid	X-ous acid
	ex. Br^{-} bromide HBr hydrobromic acid	ex. $\text{C}_2\text{H}_3\text{O}_2^{-}$ acetate $\text{HC}_2\text{H}_3\text{O}_2$ acetic acid	ex. IO_3^{-} iodate HIO_3 iodic acid
	ex. CN^{-} cyanide HCN hydrocyanic acid	ex. PO_4^{3-} phosphate H_3PO_4 phosphoric acid	ex. SO_3^{2-} sulfite H_2SO_3 sulfurous acid

Ions And Ionic Compounds Study Guide

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Table of Contents Ions And Ionic Compounds Study Guide

1. Understanding the eBook Ions And Ionic Compounds Study Guide
 - The Rise of Digital Reading Ions And Ionic Compounds Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Ions And Ionic Compounds Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ions And Ionic Compounds Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ions And Ionic Compounds Study Guide

- Personalized Recommendations
- Ions And Ionic Compounds Study Guide User Reviews and Ratings
- Ions And Ionic Compounds Study Guide and Bestseller Lists
- 5. Accessing Ions And Ionic Compounds Study Guide Free and Paid eBooks
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 - Ions And Ionic Compounds Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ions And Ionic Compounds Study Guide
 - Highlighting and Note-Taking Ions And Ionic Compounds Study Guide
 - Interactive Elements Ions And Ionic Compounds Study Guide
- 8. Staying Engaged with Ions And Ionic Compounds Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ions And Ionic Compounds Study Guide
- 9. Balancing eBooks and Physical Books Ions And Ionic Compounds Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ions And Ionic Compounds Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ions And Ionic Compounds Study Guide
 - Setting Reading Goals Ions And Ionic Compounds Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ions And Ionic Compounds Study Guide

- Fact-Checking eBook Content of Ions And Ionic Compounds Study Guide
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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programs chapter 02 chapter 03 chapter 04 chapter 05 chapter 06 chapter 07 chapter 08

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