

## WORKSHEET: BALANCING EQUATIONS

Balance the following equations by placing small whole-number coefficients in the appropriate places.

### Synthesis

1.  $\text{Ba} + \text{O}_2 \rightarrow \text{BaO}$
2.  $\text{C} + \text{S} \rightarrow \text{CS}_2$
3.  $\text{Li} + \text{O}_2 \rightarrow \text{Li}_2\text{O}$
4.  $\text{Mg} + \text{N}_2 \rightarrow \text{Mg}_3\text{N}_2$
5.  $\text{FeCl}_2 + \text{Cl}_2 \rightarrow \text{FeCl}_3$

### Decomposition or Analysis

6.  $\text{KClO}_3 \rightarrow \text{KCl} + \text{O}_2$
7.  $\text{Ag}_2\text{O} \rightarrow \text{Ag} + \text{O}_2$
8.  $\text{CuCO}_3 \rightarrow \text{CuO} + \text{CO}_2$
9.  $\text{AuBr}_3 \rightarrow \text{Au} + \text{Br}_2$
10.  $\text{UF}_4 \rightarrow \text{U} + \text{F}_2$

### Single Replacement

11.  $\text{Zn} + \text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$
12.  $\text{Zn} + \text{CuSO}_4 \rightarrow \text{ZnSO}_4 + \text{Cu}$
13.  $\text{Cu} + \text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + \text{Ag}$
14.  $\text{K} + \text{H}_2\text{O} \rightarrow \text{KOH} + \text{H}_2$
15.  $\text{Al} + \text{CuCl}_2 \rightarrow \text{AlCl}_3 + \text{Cu}$

### Double Replacement

16.  $\text{BaCl}_2 + \text{Na}_2\text{SO}_4 \rightarrow \text{NaCl} + \text{BaSO}_4$
17.  $\text{ZnCl}_2 + (\text{NH}_4)_2\text{S} \rightarrow \text{NH}_4\text{Cl} + \text{ZnS}$
18.  $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
19.  $\text{FeS} + \text{HCl} \rightarrow \text{FeCl}_2 + \text{H}_2\text{S}$
20.  $\text{AlCl}_3 + \text{NaOH} \rightarrow \text{Al}(\text{OH})_3 + \text{NaCl}$

### Combustion

21.  $\text{CH}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
22.  $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
23.  $\text{C}_3\text{H}_8 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
24.  $\text{C}_4\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
25.  $\text{C}_5\text{H}_{12} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

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