

# How to Name Chemical Compounds

is it ionic? is an anion,  
or does it contain a polyatomic ion (as provided by H<sup>+</sup>)?

## Ionic compound

[cation's name] [anion's name]

main- element	<ul style="list-style-type: none"> <li>take its element name!</li> <li>ex. <math>\text{Al}^{3+}</math>: aluminum</li> <li>if transition metal*, include charge in Roman numeral</li> <li>ex. <math>\text{Cr}^{2+}</math>: chromium (II)</li> </ul>	<ul style="list-style-type: none"> <li>get element's name, chop end, add "-ide"</li> <li>ex. <math>\text{F}^{-}</math>: fluorine → fluoride ✓</li> </ul>
poly- element	<ul style="list-style-type: none"> <li>just use its name!</li> <li>ex. <math>\text{NH}_4^{+}</math>: ammonium</li> </ul>	<ul style="list-style-type: none"> <li>just use its name!</li> <li>ex. <math>\text{CO}_3^{2-}</math>: carbonate</li> </ul>

\* exceptions (no Roman numerals):

silver ion: always  $\text{Ag}^{+}$

zinc ion: always  $\text{Zn}^{2+}$

cadmium ion: always  $\text{Cd}^{2+}$

## Molecular compound

is it  $\text{HCO}_3^{-}$ ? otherwise ...

Acid! name and formula

[#]-[1<sup>st</sup> element name] [#]-[2<sup>nd</sup> element root]-"ide"

↑ skip if "mono-" ↑ skip if 1<sup>st</sup> element is H

### Numerical prefixes

- |             |             |
|-------------|-------------|
| • 1: mono-  | • 6: hexa-  |
| • 2: di-    | • 7: hepta- |
| • 3: tri-   | • 8: octa-  |
| • 4: tetra- | • 9: nona-  |
| • 5: penta- | • 10: deca- |

for  
first element: never  
ex.  $\text{N}_2\text{O}_5$ : dinitrogen  
pentoxide

### Acids

Ignore all  $\text{H}^{+}$ , and look at the name of the remaining anion.

Anion	X-ide	X-ate	X-ite
Acid	hydro-X-ic acid	X-ic acid	X-ous acid
	ex. $\text{Br}^{-}$ bromide $\text{HBr}$ hydrobromic acid	ex. $\text{C}_2\text{H}_3\text{O}_2^{-}$ acetate $\text{HC}_2\text{H}_3\text{O}_2$ acetic acid	ex. $\text{IO}_3^{-}$ iodate $\text{HIO}_3$ iodic acid
	ex. $\text{CN}^{-}$ cyanide $\text{HCN}$ hydrocyanic acid	ex. $\text{PO}_4^{3-}$ phosphate $\text{H}_3\text{PO}_4$ phosphoric acid	ex. $\text{SO}_3^{2-}$ sulfite $\text{H}_2\text{SO}_3$ sulfurous acid

# Naming Molecular Compounds Study Guide

**RS Peters**



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