

1 Generally transition elements and their salts are coloured due to the presence of unpaired electrons in metal ions. Which of the following compounds are coloured?

- KMnO_4
- $\text{Ce}(\text{SO}_4)_2$
- TiCl_3
- Cu_2Cl_2

Choose the correct option

- (a, b)
- (b, c)
- (c, d)
- (a, d)

2 General electronic configuration of actinoids is $(n-2)f^{1-14}(n-1)d^{0-2}ns^2$. Which of the following actinoids have one electron in 6d orbital?

- U (Atomic no. 92)
- Np (Atomic no. 93)
- Pu (Atomic no. 94)
- Am (Atomic no. 95)

Choose the correct option

- (a, b)
- (b, c)
- (c, d)
- (a, d)

3 Which of the following lanthanoids show +2 oxidation state besides the characteristic oxidation state +3 of lanthanoids?

- Ce
- Eu
- Yb
- Ho

Choose the correct option

- (a, b)
- (b, c)
- (c, d)
- (a, d)

4 On heating lead nitrate forms oxides of nitrogen and lead. The oxides formed are-

- N_2O , PbO
- NO_2 , PbO
- NO , PbO
- NO , PbO_2

5 A brown ring is formed in the ring test for NO_3^- ion. It is due to the formation of-

- $[\text{Fe}(\text{H}_2\text{O})_5(\text{NO})]^{2+}$
- FeSO_4 , NO_2
- $[\text{Fe}(\text{H}_2\text{O})_4(\text{NO})_2]^{2+}$
- FeSO_4 , HNO_3

6 The oxidation state of the central atom in the anion of the compound NaH_2PO_2 will be-

- +3
- +5
- +1
- 3

7 A black compound of manganese reacts with a halogen acid to give greenish-yellow gas. When an excess of this gas reacts with NH_3 an unstable trihalide is formed. In this process, the oxidation state of nitrogen changes from-

- 3 to +3
- 3 to 0
- 3 to +5
- 0 to -3

8 Which of the following is isoelectronic pair?

- ICl_2 , ClO_2
- BrO_2^- , BrF_2^+
- ClO_2 , BrF
- CN^- , O_3

9 Which of the following statements are correct for SO_2 gas?

- It acts as a bleaching agent in moist conditions.
 - Its molecule has linear geometry.
 - Its dilute solution is used as a disinfectant.
 - It can be prepared by the reaction of dilute H_2SO_4 with metal sulphide.
- (a, b)
 - (b, c)
 - (c, d)
 - (a, c)

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J. J. Burton



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