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Meatloaf Stew Recipe

CO Houle



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web in today's experiment you will determine the solubility product K_{sp} of calcium hydroxide $Ca(OH)_2$ by measuring the concentration of $Ca(OH)_2$ in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the K_{sp} of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of $Ca(OH)_2$ is represented in equation form as shown below $Ca(OH)_2(s) \rightleftharpoons Ca^{2+}(aq) + 2OH^{-}(aq)$ the solubility product expression describes in mathematical terms the equilibrium

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web K_{sp} for calcium hydroxide calcium hydroxide is a soft white caustic powder used in making mortar cements calcium salts paints and petrochemicals it is also used in saltwater aquaria to make up kalkwasser limewater solutions for reef tanks and is used as a pH regulating agent notice that calcium hydroxide is divalent and thus has twice

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web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the solution $Ca(OH)_2 \rightarrow Ca^{2+} + 2OH^{-}$ $5 \times 10^{-2} M$ calcium ions $5 \times 10^{-2} M$ $0 M$ hydroxide ions calculate calcium hydroxide K_{sp} using the concentrations of calcium ions and hydroxide ions

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web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the K_{sp} of a slightly soluble basic salt and how solubility changes with temperature K_{sp} is another form of equilibrium expression that deals with solubility the reaction that will take place is $Ca(OH)_2 \rightleftharpoons Ca^{2+} + 2OH^{-}$

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web question experiment 8 determining K_{sp} for calcium hydroxide post lab report sheet data table 2 points trial 1 trial 2 HCl molarity $0.050 M$ volume of filtered $Ca(OH)_2$ solution ml 5.00 5.00 3.71 13.75 equivalence point volume ml OH^{-} in saturated $Ca(OH)_2$ M Ca^{2+} in saturated $Ca(OH)_2$ M K_{sp} of $Ca(OH)_2$ average K_{sp} of $Ca(OH)_2$

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web the chemical equation used was $Ca(OH)_2(s) \rightleftharpoons Ca^{2+}(aq) + 2OH^{-}(aq)$ the solubility product was determined to be $K_{sp} Ca^{2+} OH^{-}$ from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium

hydroxide was filtered before starting the titration

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web jul 12 2023 because the concentration of a pure solid such as $\text{Ca}_3(\text{PO}_4)_2$ is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore $K_{\text{sp}} = [\text{Ca}^{2+}]^3 [\text{PO}_4^{3-}]^2$

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web lab report 8 determining the ksp of calcium hydroxide the purpose of this was to titrate a saturated $\text{Ca}(\text{OH})_2$ solution with a standard HCl solution and find the ksp or the OH^- concentration was found and then the concentration was found and used for this experiment a titration of calcium hydroxide with hydrochloric acid

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web therefore it was determined through error calculations and ease of use that the most reliable valid method to use to get the closest ksp value to the theoretical ksp for $\text{Ca}(\text{OH})_2$ was the determination through pH method 2 introduction calcium hydroxide is a strong base soluble in water

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web calculate the ksp for calcium hydroxide Ca^{2+} OH^- 2.0 $\times 10^{-2}$ ksp 5×10^{-4} find the accepted value of the ksp for calcium hydroxide and compare it with your value discuss the discrepancy and suggest possible sources of experimental error accepted value for ksp for sodium hydroxide is 5×10^{-6}

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