

MS Word Exercise

1. Locate Microsoft Word and open a new document.
2. Type the following text:

He was an old man who fished alone in a skiff in the Gulf Stream and he had gone eighty-four days now without taking a fish.

Ernest Hemingway: The Old Man and the Sea (1952)

3. Italicize the text.
4. Underline the title of the story.
5. Bold the author's name.
6. Increase the font size to 14.
7. Change the font style to (your choice!).
8. Adjust the line spacing to 2.0 (double space).
9. Spellcheck your work.
10. Use Save As to save your document to the M Drive. Name the document The Old Man and the Sea.
11. Insert a picture onto your document (find an image from the internet, save it to the M drive).
12. Add a numbered list with your top three favourite authors (or movie stars, etc.).
13. Save your work.

Ms Word Exercises For Practice

RJ Alexander



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Ms Word Exercises For Practice Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has been apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Ms Word Exercises For Practice**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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web google classroom for the following redox equation $\text{mno}_4^- + \text{a}^- + \text{aq} \rightarrow \text{so}_4^{2-} + \text{g} + \text{mn}^{2+} + \text{aq} + \text{hso}_4^- + \text{a}^- + \text{aq}$ what will be the stoichiometric coefficient of hso_4^- if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web solution 1 the two half reactions balanced as if in acidic solution $2\text{cn}^- + \text{au} \rightarrow \text{au}^- + \text{cn}_2 + 2\text{e}^-$ $2\text{h}_2\text{o}_2 \rightarrow 2\text{h}_2\text{o} + \text{o}_2 + 2\text{e}^-$ make electrons equal convert to basic solution $4\text{cn}^- + 2\text{au} \rightarrow 2\text{au}^- + 2\text{cn}_2 + 2\text{e}^-$ multiplied by a factor of 2 $2\text{h}_2\text{o} + \text{o}_2 \rightarrow 2\text{h}_2\text{o} + 2\text{oh}^- + 2\text{e}^-$ the final answer

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation ni s ni a 2 a q 2 e now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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