

## EQUILIBRIUM TEST

### MULTIPLE CHOICE QUESTIONS

8) Consider the following reaction:



Which one of the following statements is correct?

- A. Since very large, the reaction is irreversible.
- B. Changing the pressure will affect the value of  $K_c$ .
- C. It is an very large, the reaction is effectively irreversible.
- D. Changing the temperature will not affect the value of  $K_c$ .

9) The following reaction, in which nitrogen trioxide,  $\text{N}_2\text{O}_3$ , is converted to nitrogen dioxide,  $\text{NO}_2$ , occurs at equilibrium:



In a given closed system, the equilibrium constant for this reaction is 0.50. If  $K_c$  and the initial concentration of  $\text{N}_2\text{O}_3$  at equilibrium is 0.5000 M, at this temperature, the initial concentration of  $\text{NO}_2$  at equilibrium is

- A. 0.5000 M
- B. 1.000 M
- C. 1.50 M
- D. 2.00 M

10) Chloroacetic acid,  $\text{ClCH}_2\text{COOH}$ , reacts with water,  $\text{H}_2\text{O}$ , to form methanoic acid,  $\text{H}_2\text{COOH}$ ,  $\text{H}_2\text{COO}^-$ , chloroacetate ions,  $\text{ClCH}_2\text{COO}^-$  and hydronium ions,  $\text{H}_3\text{O}^+$  according to the equilibrium below:



A solution that contains 0.1 M chloroacetic acid is placed in a vessel and left overnight. At that time the following reaction reaches equilibrium. As the reaction goes longer is made in the system, which one of the following statements is most likely to result in an increase in the concentration of  $\text{H}_3\text{O}^+$  ions?

- A. Consider that addition.
- B. Bubble in more  $\text{CO}_2$  gas.
- C. Add a few drops of phenolphthalein indicator.
- D. Add a few drops of 0.1 M sodium hydroxide,  $\text{NaOH}$ , solution.

# Multiple Choice Concept Questions On Equilibrium

**Greg Jacobs**



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