

# Thermochemistry

## Multiple Choice Questions

- B** 1. When the three types of energy changes are arranged in order of decreasing molar heat values, what is the correct order?  
**A.** chemical, phase, nuclear  
**B.** nuclear, chemical, phase  
**C.** nuclear, phase, chemical  
**D.** phase, chemical, nuclear
- B** 2. The temperature remains constant as energy is added to a substance. How may the substance be changing?  
**A.** from a gas to a solid  
**B.** from a liquid to a gas  
**C.** from a liquid to a solid  
**D.** in the amount of kinetic energy
- A** 3. Which metal requires the most energy to raise 1.00 g of it by 1.00°C?  
**A.** aluminum  
**B.** copper  
**C.** lead  
**D.** nickel
- B** 4. A substance has a molar heat of combustion of -850.4 kJ/mol. When 0.285 mol of the substance is burned in a calorimeter containing 8.60 kg of water, what is the increase of the water temperature?  
**A.** 0.156°C  
**B.** 6.41°C  
**C.** 7.89°C  
**D.** 12.8°C
- B** 5. The addition of 9.54 kJ of heat is required to raise the temperature of 225.0 g of a liquid hydrocarbon from 20.5°C to 45.0°C. What is the heat capacity of this hydrocarbon?  
**A.** 0.94 J/g·°C  
**B.** 1.73 J/g·°C  
**C.** 1.88 J/g·°C  
**D.** 9.42 J/g·°C
- C** 6. If the molar heat of fusion of sodium is 2.63 kJ/mol, how much energy is needed to melt 180.0 g of solid sodium at its melting point?  
**A.** 2.63 kJ  
**B.** 2.93 kJ  
**C.** 20.6 kJ  
**D.** 473 kJ
- B** 7. Which process is exothermic?  
**A.** the boiling of liquid nitrogen  
**B.** the freezing of water  
**C.** the sublimation of dry ice  
**D.** the vaporization of water
- A** 8. If the heat of formation for the chemical reaction,  $\text{S}_8(\text{s}) + 12\text{O}_2(\text{g}) \rightarrow 8\text{SO}_2(\text{g})$ , is -395.7 kJ/mol of  $\text{SO}_2$ , how can the equation be rewritten to include the heat term?  
**A.**  $\text{S}_8(\text{s}) + 12\text{O}_2(\text{g}) \rightarrow 8\text{SO}_2(\text{g}) + 3166 \text{ kJ}$   
**B.**  $\text{S}_8(\text{s}) + 12\text{O}_2(\text{g}) \rightarrow 8\text{SO}_2(\text{g}) - 395.7 \text{ kJ}$   
**C.**  $\text{S}_8(\text{s}) + 12\text{O}_2(\text{g}) \rightarrow 8\text{SO}_2(\text{g}) + 395.7 \text{ kJ}$   
**D.**  $\text{S}_8(\text{s}) + 12\text{O}_2(\text{g}) + 3166 \text{ kJ} \rightarrow 8\text{SO}_2(\text{g})$
- A** 9. Which statement correctly describes the energy changes that take place when a solid changes to a liquid at constant temperature?  
**A.** The potential energy increases, and the kinetic energy remains constant.  
**B.** The potential energy remains constant, and the kinetic energy increases.  
**C.** The potential energy decreases, and the kinetic energy decreases.  
**D.** The potential energy increases, and the kinetic energy decreases.
- B** 10. Why does warm water have a more rapid rate of evaporation than cold water?  
**A.** It has a higher viscosity.  
**B.** Its molecules have a higher average kinetic energy.  
**C.** Its molecules have more attraction for one another.  
**D.** More of its molecules have nearly the same kinetic energy.
- B** 11. If the heat of fusion of a substance is 20 kJ/mol, what heat is released when 1.0 mol of liquid at the melting point freezes?  
**A.** 10 kJ  
**B.** 20 kJ  
**C.** 40 kJ  
**D.** 60 kJ
- C** 12. The amount of 2.00 mol of a solid ionic compound was dissolved in 225 mL of water in an insulated container. The initial temperature of the water was 20.0°C and the final temperature of the water was 47.8°C. What is  $\Delta H_{\text{soln}}$ ?  
**A.** -27.8 kJ/mol  
**B.** -26.2 kJ/mol  
**C.** -13.1 kJ/mol  
**D.** -6.26 kJ/mol
- D** 13. What does it mean if the  $\Delta H$  value for a chemical reaction is positive?  
**A.** kinetic energy is increasing in the system  
**B.** potential energy is decreasing in the system  
**C.** products have less potential energy than reactants  
**D.** reactants have less potential energy than products.
- A** 14. Living plants produce glucose in the process of photosynthesis according to this equation:  
 $6\text{CO}_2(\text{g}) + 6\text{H}_2\text{O}(\text{l}) + \text{energy} \rightarrow \text{C}_6\text{H}_{12}\text{O}_6(\text{s}) + 6\text{O}_2(\text{g})$   
 Is this reaction endothermic or exothermic, and is the value of  $\Delta H^\circ$  positive or negative?  
**A.** endothermic, positive  
**B.** endothermic, negative  
**C.** exothermic, positive  
**D.** exothermic, negative

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**Jianjun Gao**



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