

Name: _____

Answer Key: Simulation: Ionic and Covalent Bonding

Background

In this investigation you will bond select atoms. Based upon the types of atoms that you choose to combine, you will create either an ionic compound or a covalent compound. You will have the opportunity to analyze the differences between these different types of compounds and to predict the number of atoms needed to create each, as well as learn how to appropriately name them.

1. Describe the difference between an atom and a molecule:

Atoms are the smallest single unit of a particular element and are the smallest functional unit of matter; atoms are neutral, and each element has its own unique type of atom. Molecules are neutral groups of more than one atom stuck together with a covalent bond (sharing electrons); they can be more than one atom of the same element or of different elements.

2. Where are metal atoms located on the periodic table? Where are non-metal atoms located on the periodic table?

Metals are on the left and center of the periodic table and they make up about 80% of the elements, so they take up the most space on the periodic table. Nonmetals are in the upper right corner of the periodic table, as well as hydrogen, the very first element on the upper left. They only make up about 15% of the elements.

3. What subatomic particle(s) participate in chemical bonding?

Electrons are the subatomic particles that participate in chemical bonding.

4. In your own words, define *valence electron*.

Valence electrons are the outermost electrons on an atom, and they are the ones that form an atom's bonds, since they are on the outside and accessible to other atoms that might form bonds. They can be exchanged between atoms or shared

5. How can you determine the number of valence electrons in an atom using the periodic table?

The group numbers can help determine the number of valence electrons. Group 1A elements have 1 valence electron, 2A have 2, 3A have 3, etc. through 8 A. This skips over the transition metals (the elements in the "B" groups in the middle section of the periodic table), as their valence electrons are not as easily determined. If there are no letters in your periodic table group numbers and they just go from 1-18, subtract 10 from groups 13-18 and that is the number of valence electrons. (Again, this skips the transition metals.)

6. Draw a Lewis Dot Structure for the following atoms:

a. Strontium (Sr)



b. Carbon (C)



c. Iodine (I)



d. Xenon (Xe)



*Check your answers before moving on to the next portion of the activity.

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R Sanford



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