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Robert R. Redfield



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with 1200 ml of 2M HNO_3 if the final density of solution is same as pure water then molarity of the iron in resulting solution which decides the nature of the above solution is 1 0.25M 2 0.5M 3 0.888M 4 none of

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$$m = \frac{\text{moles of solute}}{\text{kg of solvent}}$$

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solution of molarity 1.5 M requires 1.5 mol of Na to every litre of solvent

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$$M_{mix} = \frac{M_1 V_1 + M_2 V_2 + \dots + M_n V_n}{V_1 + V_2 + \dots + V_n}$$
 with increase in temperature which of these changes a molality b fraction of solute present in water

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