

A Microscale Titration

Propose a series of microscale titrations involving the formation of a salt between a weak acid and a weak base. Record the observed color change.

Observe color changes.

1. Propose a series of microscale titrations between a:

Weak acid: acetic acid

Weak base: ammonia

Observe 20 drops:

2. After about 5, the color of the solution is:

yellow to orange

3. Repeat the titration from Procedures 1, 2, and 3, using:

Weak acid: acetic

Weak base: ammonia

Observe 20 drops:

4. What do you think is the color of the average compound that is formed that was used?

Observed color: yellow to orange

5. What does the color of the solution suggest the identity of the acid, and the compound that forms (molecular formula) molecular structure of the acid and the compound?

Equation of acid-base reaction that observed: CH₃COOH + NH₃

6. What does the color of the solution suggest the identity of the acid-base compound?

Equation of acid-base reaction: CH₃COOH + NH₃

Microscale Titration Lab Answers

JS Bruner



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