

Procedure for making NaCl-Aid (for every lab group of 3-4 people):

1. Using your calculations from question #4 above, measure out the correct amount of solid NaCl-Aid powder (grams) into the assigned cup to make a 0.1 M solution. **Be careful not to include the mass of the cup in your measurement!**
2. Transfer the powder into the assigned "mixing cup".
3. Add water into the "mixing cup" until you have 100 mL of solution. Stir with a spoon or stick.
4. Observe and taste the solutions you have made. You can have one "designated taster" or you can pour a little into separate little cups for each group member to taste. Record in data table.
5. Repeat steps 1-3 in order to make the 0.4 M and 0.7 M solutions.
6. Clean-Up... Throw away cups that you drank out of, but save the mixing & measuring cups.

Data / Observations:

	0.1 M solution	0.4 M solution	0.7 M solution
Observations (taste (color), smell, notes)	Answers will vary	Answers will vary	Answers will vary

Calculations / Analysis:

1. Which concentration that you tasted was closest to the ideal concentration of NaCl-Aid? What was wrong with each of the other solutions that you made?

Answers will vary... Some solutions will not have enough NaCl-Aid for your taste, while others will have too much.

2. What was the solute used in this lab? What was the solvent?

Solute = NaCl-Aid powder Solvent = water

3. What is molarity?

Molarity is a measure of the number of moles of solute per liter of solution.

Questions for Discussion:... You must show the formula, all your work, and **highlight** your answer!

1. Write the formula for the 4 different ways to calculate concentrations:

- a. Molarity

$$M = \frac{\text{moles}}{\text{volume}}$$

- a. grams/Liter

$$\frac{\text{grams of solute}}{\text{volume of solution}}$$

- a. % composition

$$\frac{\text{grams of solute}}{\text{grams of solution}} \times 100$$

- a. Parts per million

$$\frac{\text{grams of solute}}{\text{grams of solution}} \times 10^6$$

Determine the following concentrations, using the 4 different methods:

****Disclaimer: 1 gram NaCl = 1 mL of 1% = 1000 mL****

2. If 1.80 moles of NaCl was dissolved in enough water to make 0.60 L of solution:

- a. Molarity?

$$M = \frac{1.8 \text{ mol}}{0.6 \text{ L}}$$

$$M = 3.000 \text{ M NaCl}$$

- a. Grams/Liter?

$$\frac{180 \text{ g NaCl}}{0.6 \text{ L}}$$

$$= 300 \text{ g/L}$$

- a. % composition?

$$\frac{180 \text{ g NaCl}}{1800 \text{ g}} \times 100$$

$$= 10.00\% \text{ NaCl}$$

- a. Parts per million?

$$\frac{180 \text{ g NaCl}}{1800 \text{ g}} \times 10^6$$

$$= 100,000 \text{ ppm}$$

3. If 20 grams of NaOH was dissolved in enough water to make a 250 mL NaOH solution:

- a. Molarity?

$$M = \frac{0.5 \text{ mol}}{0.25 \text{ L}}$$

$$M = 2.0 \text{ M NaOH}$$

- a. Grams/Liter?

$$\frac{20 \text{ g NaOH}}{0.25 \text{ L}}$$

$$= 80 \text{ g/L}$$

- a. % composition?

$$\frac{20 \text{ g NaOH}}{250 \text{ g}} \times 100$$

$$= 8\% \text{ NaOH}$$

- a. Parts per million?

$$\frac{20 \text{ g NaOH}}{250 \text{ g}} \times 10^6$$

$$= 80,000 \text{ ppm}$$

Kool Aid Lab Answer Key

Jiyuan Zhang

A red circular graphic with a gradient, appearing as a stylized arrow or a light effect, positioned to the right of the name Jiyuan Zhang.

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the rest of his life eventually taking over the photography business of kahlo s mother s family

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web kahlo stated that she was born at the family home la casa azul the blue house but according to the official birth registry the birth took place at the nearby home of her maternal grandmother kahlo s parents were photographer guillermo kahlo 1871 1941 and matilde calderón y gonzález 1876 1932 and they were thirty six and thirty

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