# Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- a. Write the balanced chemical equation.
- b. Identify the given (with units) and what you want to find (with units).
- c. Show set up with units. Give final answer with units and label.
- 1. Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?

Before 
$$2 H_2 O \rightarrow 2 H_2 + O_2$$

Before  $1.99 \text{ mol} \rightarrow 1.99 \text{ mol} \rightarrow 1.00 \text{ mol}$ 

Change  $-2.99 \text{ mol} \rightarrow 1.00 \text{ mol}$ 

After  $0 \text{ mol} \rightarrow 2.00 \text{ mol}$ 
 $1.00 \text{ mol} \rightarrow 2.00 \text{ mol$ 

2. Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?

# **Modeling Chemistry U8 V2 Answers**

**KJ Lindholm-Leary** 

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