

Chemistry – Unit 5 Worksheet 2

1. An old (pre-1987) penny is nearly pure copper. If such a penny has a mass of 3.3 g, how many moles of copper atoms would be in one penny?

$$3.3 \text{ g} \times \frac{1 \text{ mole}}{63.5 \text{ g}} = 5.2 \times 10^{-2} \text{ mol or } 0.052 \text{ mol}$$

2. Four nails have a total mass of 4.42 grams. How many moles of iron atoms do they contain?

$$4.42 \text{ g} \times \frac{1 \text{ mole}}{55.8 \text{ g}} = 7.92 \times 10^{-2} \text{ mol or } 0.0792 \text{ mol}$$

3. A raindrop has a mass of 0.050 g. How many moles of water does a raindrop contain?

$$0.050 \text{ g} \times \frac{1 \text{ mole}}{18.0 \text{ g}} = 2.78 \times 10^{-3} \text{ mol} \rightarrow 2.8 \times 10^{-3} \text{ mol}$$

4. What mass of water would you need to have 15.0 moles of H₂O?

$$15.0 \text{ moles} \times \frac{18.0 \text{ g}}{1 \text{ mole}} = 2.70 \times 10^2 \text{ g}$$

5. One box of Morton's Salt contains 737 grams. How many moles of sodium chloride is this?

$$737 \text{ g} \times \frac{1 \text{ mole}}{58.5 \text{ g}} = 12.6 \text{ mol}$$

Na	23.0
Cl	35.5
	<hr/> 58.5 g

6. A chocolate chip cookie recipe calls for 0.050 moles of baking soda (sodium bicarbonate). How many grams should the chef mass out?

$$0.050 \text{ moles} \times \frac{84.0 \text{ g}}{1 \text{ mole}} = 4.2 \text{ g}$$

Na	23.0
H	1.0
C	12.0
O	48.0
	<hr/> 84.0

7. Rust is iron(III) oxide. The owner of a 1959 Cadillac convertible wants to restore it by removing the rust with oxalic acid, but he needs to know how many moles of rust will be involved in the reaction. How many moles of iron(III) oxide are contained in 2.50 kg of rust?

$$2.50 \text{ kg} \times \frac{1000 \text{ g}}{1 \text{ kg}} \times \frac{1 \text{ mole}}{159.6 \text{ g}} = 15.7 \text{ mol Fe}_2\text{O}_3$$

Fe	111.6
O	48.0
	<hr/> 159.6

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IM Harris



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Table of Contents Modeling Chemistry Unit 1 Answers Key

1. Understanding the eBook Modeling Chemistry Unit 1 Answers Key
 - The Rise of Digital Reading Modeling Chemistry Unit 1 Answers Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry Unit 1 Answers Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry Unit 1 Answers Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry Unit 1 Answers Key
 - Personalized Recommendations
 - Modeling Chemistry Unit 1 Answers Key User Reviews and Ratings
 - Modeling Chemistry Unit 1 Answers Key and Bestseller Lists
5. Accessing Modeling Chemistry Unit 1 Answers Key Free and Paid eBooks
 - Modeling Chemistry Unit 1 Answers Key Public Domain eBooks
 - Modeling Chemistry Unit 1 Answers Key eBook Subscription Services

- Modeling Chemistry Unit 1 Answers Key Budget-Friendly Options
- 6. Navigating Modeling Chemistry Unit 1 Answers Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Modeling Chemistry Unit 1 Answers Key Compatibility with Devices
 - Modeling Chemistry Unit 1 Answers Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Chemistry Unit 1 Answers Key
 - Highlighting and Note-Taking Modeling Chemistry Unit 1 Answers Key
 - Interactive Elements Modeling Chemistry Unit 1 Answers Key
- 8. Staying Engaged with Modeling Chemistry Unit 1 Answers Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry Unit 1 Answers Key
- 9. Balancing eBooks and Physical Books Modeling Chemistry Unit 1 Answers Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry Unit 1 Answers Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Modeling Chemistry Unit 1 Answers Key
 - Setting Reading Goals Modeling Chemistry Unit 1 Answers Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modeling Chemistry Unit 1 Answers Key
 - Fact-Checking eBook Content of Modeling Chemistry Unit 1 Answers Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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