

## Unit 6 – Thermodynamics (7-9%)

Changes in a substance's properties or change into a different substance requires an exchange of energy.

**6.1 Endothermic and Exothermic Processes:** Explain the relationship between experimental observations and energy changes associated with a chemical or physical transformation.

- Temperature changes in a system indicate energy changes.
- Energy changes in a system can be described as endothermic and exothermic processes such as the heating or cooling of a substance, phase changes, or chemical transformations.
- When a chemical reaction occurs, the energy of the system either decreases (exothermic reaction), increases (endothermic reaction), or remains the same. For exothermic reactions, the energy lost by the reacting species (system) is gained by the surroundings, as heat transfer from or work done by the system. Likewise, for endothermic reactions, the system gains energy from the surroundings by heat transfer to or work done on the system.
- The formation of a solution may be an exothermic or endothermic process, depending on the relative strengths of intermolecular/interparticle interactions before and after the dissolution process.

**6.2 Energy Diagrams:** Represent a chemical or physical transformation with an energy diagram.

- A chemical or physical process is labeled with an energy diagram that shows the endothermic or exothermic nature of the process.

# Unit 6 Objectives

## Thermodynamics

- The particles in a warmer body have a greater average kinetic energy than those in a cooler body.
- Collisions between particles in a warmer body and a cooler body are called exothermic processes.
- Eventually, thermal equilibrium is reached as the particles continue to collide. At thermal equilibrium, the average kinetic energy of both bodies is the same and there is no net energy transfer.

AP Chemistry

REFERENCE



# Modeling Chemistry Unit 6 Objectives

**Minjie Lin**



**Modeling Chemistry Unit 6 Objectives:**

## Decoding **Modeling Chemistry Unit 6 Objectives**: Revealing the Captivating Potential of Verbal Expression

In a time characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its ability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Modeling Chemistry Unit 6 Objectives**," a mesmerizing literary creation penned with a celebrated wordsmith, readers set about an enlightening odyssey, unraveling the intricate significance of language and its enduring impact on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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