

Chemistry – Unit 5 Worksheet 1

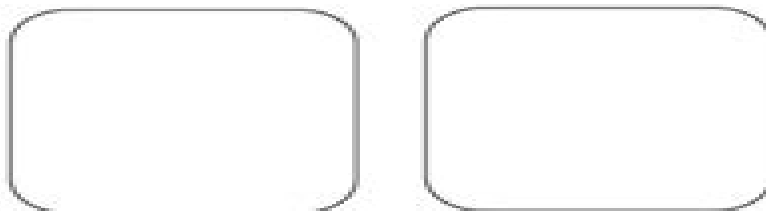
Molar Masses of the Elements

Relative Mass From Gases

We have established that the combining ratio of gases can be explained if two assumptions are made:

1. Equal volumes of gases contain the same number of molecules at the same pressure and temperature.
2. Some pure elemental gases are clustered into pairs to form diatomic molecules.

Use these assumptions and particle diagrams to explain the fact that the density of oxygen gas at standard temperature and pressure is 1.43 g/liter, whereas the density of hydrogen gas under these conditions is 0.089 g/liter. How many times more massive is one molecule of oxygen than one molecule of hydrogen?



You shouldn't conclude that chemists were able to determine the molar masses of all the elements using this technique. Measurements of the density of the gaseous phase of many of the elements are difficult, if not impossible. However, we are going to see that chemists could use another tool – the *percent composition of compounds* to determine molar masses.

Relative Mass From Compounds

Many substances combine with oxygen to form a type of compound called an oxide. In the previous unit we saw that such combinations often occur in multiple proportions. John Dalton made the assumption that the lowest ratio was a 1:1 combination of elements. For now, we will make a similar assumption. *We may have to re-examine this assumption later.*

In Unit 4 you examined % composition data of several compounds. For example, 100 g of the simplest oxide of carbon contains 42.9g of C and 57.1 g of O. To be able to compare masses of elements to the mass of oxygen, it is important to use the same mass of oxygen in all of the ratios. Using a proportion, one can show that 75.0 g of C combines with 100 g of oxygen.

$$\frac{42.9\text{gC}}{57.1\text{gO}} = \frac{75.0\text{gC}}{100\text{gO}}$$

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Chemistry in the Community (ChemCom) American Chemical Society, 2011-06-17 Touted as the most successful NSF funded project published *Chemistry in the Community ChemCom* by the American Chemical Society ACS offers a meaningful and memorable chemistry program for all levels of high school students ChemCom covers traditional chemistry topics within the context of societal issues and real world scenarios Centered on decision making activities where students are responsible for generating data in an investigating analyzing that data and then applying their chemistry knowledge to solve the presented problem The text is intensively laboratory based with all 39 of the investigations integrated within the text not separate from the reading With the ChemCom program students learn more organic and biochemistry more environmental and industrial chemistry and more on the particulate nature of matter than other textbooks all within the relevance of solving problems that arise in everyday life Meticulously updated to meet the needs of today's teachers and students the new sixth edition of ChemCom adheres to the new science framework as well as the forthcoming next generation of science standards Incorporating advances in learning and cognitive sciences ChemCom's wide ranging coverage builds upon the concepts and principles found in the National Science Education Standards Correlations are available showing how closely aligned ChemCom is to these and other state standards

Handbook of CALL Teacher Education and Professional Development Dara Tafazoli, Michelle Picard, 2023-04-26 This comprehensive handbook provides an overview of current trends in computer assisted language learning CALL teacher education and professional development across the globe It highlights theories and practices in CALL teacher education and professional development in five sections such as English language teaching including pre service teachers in service teachers teacher educators material developers course designers and researchers It explores the role of CALL teacher education and professional development in many underexplored countries such as Africa Asia Eastern Europe and the Middle East It stresses the critical role of professional development programs from the use of technology in its generic sense The theoretical and empirical chapters in the book provide a more inclusive and comprehensive picture of various aspects of CALL teacher education and professional development globally It offers context specific approaches and strategies to language teachers and teacher educators It provides pedagogical implications and suggestions for promoting digital literacy and autonomy in online education This book provides valuable insights for researchers teacher educators and teacher trainers in applied linguistics

Chemistry George Tyler Miller, 1987

Cyber Science 6 Tm' 2007 Ed. , The World of Science Sixth Grade A. Edward Ooghe, 1963

Resources in Education , 1984 New Curriculum Perspectives for Junior High School Science Paul DeHart Hurd, 1970

Research in Education , 1973 **Prentice Hall Chemistry** , 2000 **Russian Journal of Physical Chemistry** , 2001 *Inquiry Into Physical Science* Willard J. Jacobson, 1969 Adult Training , 1975 Merrill Earth Science Ralph M. Feather, Susan Leach Snyder, Dale T. Hesser, 1995 Creating the Ideal School Albert Mamary, 2007-08-13

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The Chemistry of Galaxies Marshall Lester McCall, 1982 **Encyclopedia of Soils in the Environment** Daniel Hillel, Jerry L. Hatfield, 2004. Provides a wide range of scientific knowledge on all aspects of soil science as well as the links of soils and soil science to environmental management, food production, biodiversity, climate change and many other areas of significant concern.

[Courses Catalog - University of Illinois at Urbana-Champaign](#)
University of Illinois at Urbana-Champaign, 2002. Includes undergraduate and graduate courses.

[Teaching School Physics](#)
John L. Lewis, 1972. A UNESCO source book.

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