

MATHEMATICS

JS III

SECOND TERM

THEME: ALGEBRAIC PROCESSES

SUB-THEME: ALGEBRAIC OPERATIONS

WEEK	TOPIC	PERFORMANCE OBJECTIVE	CONTENT	ACTIVITIES		TEACHING AND LEARNING RESOURCES	EVALUATION GUIDE
				TEACHER	STUDENTS		
4	Simple equations involving fractions 2	Students should be able to solve simple problems on simple equations involving fractions.	Word problems leading to simple equations involving fractions.	1. Guides students to translate word problems leading to simple equations with fractions. 2. Leads students to solve the resulting algebraic equation.	1. Translate each word problem to algebraic equation. 2. Leads students to solve the resulting algebraic equation.	Flashcards of simple squares involving fractions.	Translate given word problem into algebraic equation and solve them.
5	Simultaneous Linear Equations 1	Students should be able to: 1. compile table of values for simultaneous linear functions. 2. solve problems involving simultaneous equations in two variables graphically	1. Compilations of tables of values 2. Graphical solution of simultaneous linear equations in two variables.	1. Guides students in calculating the values of a linear equation for different values of the variables and present their solutions in a table. 2. Guides students to plot the points for the two linear equations on the same axes. 3. Leads students to locate and read points of intersection of the two graphs as the solution of the equations.	1. Computer linear equations for different values for the variable. 2. Present calculated values in tabular form. 3. Plot the points for the two linear equations on the same axes. 4. Locate and read points of intersection of the two graphs as the solution of the equation.	Table of values, Ruler, Graph board, pencil.	Students to: 1. compute table of values for given linear functions. 2. solve a given simultaneous equation in two variables graphically.

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Table of Contents Jss 1 Scheme Of Work P H E

1. Understanding the eBook Jss 1 Scheme Of Work P H E
 - The Rise of Digital Reading Jss 1 Scheme Of Work P H E
 - Advantages of eBooks Over Traditional Books
2. Identifying Jss 1 Scheme Of Work P H E
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Jss 1 Scheme Of Work P H E
 - User-Friendly Interface

-
4. Exploring eBook Recommendations from Jss 1 Scheme Of Work P H E
 - Personalized Recommendations
 - Jss 1 Scheme Of Work P H E User Reviews and Ratings
 - Jss 1 Scheme Of Work P H E and Bestseller Lists
 5. Accessing Jss 1 Scheme Of Work P H E Free and Paid eBooks
 - Jss 1 Scheme Of Work P H E Public Domain eBooks
 - Jss 1 Scheme Of Work P H E eBook Subscription Services
 - Jss 1 Scheme Of Work P H E Budget-Friendly Options
 6. Navigating Jss 1 Scheme Of Work P H E eBook Formats
 - ePub, PDF, MOBI, and More
 - Jss 1 Scheme Of Work P H E Compatibility with Devices
 - Jss 1 Scheme Of Work P H E Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Jss 1 Scheme Of Work P H E
 - Highlighting and Note-Taking Jss 1 Scheme Of Work P H E
 - Interactive Elements Jss 1 Scheme Of Work P H E
 8. Staying Engaged with Jss 1 Scheme Of Work P H E
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Jss 1 Scheme Of Work P H E
 9. Balancing eBooks and Physical Books Jss 1 Scheme Of Work P H E
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Jss 1 Scheme Of Work P H E
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Jss 1 Scheme Of Work P H E
 - Setting Reading Goals Jss 1 Scheme Of Work P H E
 - Carving Out Dedicated Reading Time

-
12. Sourcing Reliable Information of Jss 1 Scheme Of Work P H E
 - Fact-Checking eBook Content of Jss 1 Scheme Of Work P H E
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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monoxide co determine the mass of iodine I_2 which could be produced b if in the above situation only 0.160 moles of iodine I_2 was produced i what mass of iodine was produced

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$$Cu + 2 AgNO_3 \rightarrow Cu(NO_3)_2 + 2 Ag$$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } H_2S \times \frac{2 \text{ mole } H_2O}{2 \text{ mole } H_2S} = 8.3 \text{ mole } H_2O$
 $8.3 \text{ mole } H_2O \times 18 \text{ g } H_2O = 149.4 \text{ g } H_2O$ is your theoretical yield so percent yield $\frac{137.1 \text{ g } H_2O}{149.4 \text{ g } H_2O} \times 100 = 91.77\%$
 problems 6 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH_3 can be produced theoretically from the reaction of 5.0 g of n what is the limiting reagent if 8.52 g are actually formed what is the

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you have 0.216232 grams mole 21.3 grams 0.66 moles 32 g m

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web percent yield is a measure of the efficiency of the experimental design yield efficiency mass of product obtained

calculated mass of the product expected $\times 100$ in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ 4 Al 3 O_2 2 Al_2O_3 b determine the theoretical yield of Al_2O_3 160.0

g Al $\times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}$ 4 mol Al 1 mol Al_2O_3 c determine the percent yield

yield actual $\times 100$ 260.0 g $\times 100$ 86.01 theoretical 302.3 g

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web 7 Na 2 C 2 O 4 is the limiting reactant percent yield 86.6 8 only four molecules can be made 9 this amount cannot be

weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12

C is the limiting reagent 4 33 g of H 2 are left over 13

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from a set of reactants calculate how much product will be produced from the limiting reagent calculate how much reactant s

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web the percent yield is calculated as follows $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent

yield theoretical $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ based on this definition we would expect a percent yield

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CH_3OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH_3OH if 68.5 g of CO is

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