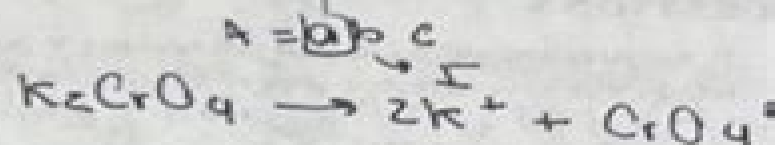


Determination of the Solubility-Product Constant for a Sparingly Soluble Salt

A. Preparation of a Calibration Curve

Initial $[\text{CrO}_4^{2-}]$ 0.0024 M

	Volume of 0.0024 M K_2CrO_4	Total volume	$[\text{CrO}_4^{2-}]$	Absorbance
1.	<u>1</u>	<u>100 mL</u>	<u>2.4×10^{-5}</u>	<u>0.072</u>
2.	<u>5</u>	<u>100 mL</u>	<u>1.2×10^{-4}</u>	<u>0.495</u>
3.	<u>10</u>	<u>100 mL</u>	<u>2.4×10^{-4}</u>	<u>1.038</u>
4.	<u>15</u>	<u>100 mL</u>	<u>3.6×10^{-4}</u>	<u>1.538</u>

$$\frac{5091.}{(2.4 \times 10^{-4})}$$

Molar extinction coefficient for $[\text{CrO}_4^{2-}]$ 1. 300 2. 4125 3. 4325 4. 4272Average molar extinction coefficient 3930.5

$$\sqrt{(300 - 3930.5)^2 + (4125 - 3930.5)^2 + (4325 - 3930.5)^2 + (4272 - 3930.5)^2}$$

Standard deviation (show calculations) 1084.6

B. Determination of the Solubility-Product Constant

$$y = 4332.1 \times 10.0$$

	Absorbance	$[\text{CrO}_4^{2-}]$	$[\text{Ag}^+]$	K_{sp} of Ag_2CrO_4
1.	<u>1.064</u>	<u>2.47×10^{-4}</u>	<u>4.94×10^{-4}</u>	<u>6.02×10^{-11}</u>
2.	<u>0.842</u>	<u>1.99×10^{-4}</u>	<u>3.96×10^{-4}</u>	<u>3.10×10^{-11}</u>
3.	<u>0.723</u>	<u>1.71×10^{-4}</u>	<u>3.41×10^{-4}</u>	<u>1.99×10^{-11}</u>

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