

# STOICHIOMETRY: MASS-MASS PROBLEMS

Name \_\_\_\_\_



How many grams of potassium chloride are produced if 25 g of potassium chlorate decompose?

$$25\text{g KClO}_3 \left( \frac{1\text{ mol}}{122.5\text{g}} \right) \left( \frac{2\text{ mol KCl}}{2\text{ mol KClO}_3} \right) \left( \frac{74.5\text{g KCl}}{1\text{ mol KCl}} \right) = 13.2\text{g KCl}$$



How many grams of hydrogen are necessary to react completely with 50.0 g of nitrogen in the above reaction?

$$50\text{g N}_2 \left( \frac{1\text{ mol N}_2}{28\text{g N}_2} \right) \left( \frac{3\text{ mol H}_2}{1\text{ mol N}_2} \right) \left( \frac{2\text{g H}_2}{1\text{ mol H}_2} \right) = 10.7\text{g H}_2$$

3. How many grams of ammonia are produced in the reaction in Problem 2?

$$50\text{g N}_2 \quad 50\text{g} + 10.7\text{g} = \boxed{60.7\text{g NH}_3}$$



How many grams of silver chloride are produced from 5.0 g of silver nitrate reacting with an excess of barium chloride?

$$5.0\text{g AgNO}_3 \left( \frac{1\text{ mol}}{170\text{g}} \right) \left( \frac{2\text{ mol AgCl}}{2\text{ mol AgNO}_3} \right) \left( \frac{143\text{g}}{1\text{ mol AgCl}} \right) = 4.21\text{g AgCl}$$

5. How much barium chloride is necessary to react with the silver nitrate in Problem 4?

$$5.0\text{g AgNO}_3 \left( \frac{1\text{ mol}}{170\text{g AgNO}_3} \right) \left( \frac{1\text{ mol BaCl}_2}{2\text{ mol AgNO}_3} \right) \left( \frac{207\text{g}}{1\text{ mol BaCl}_2} \right) = \boxed{5.0\text{g BaCl}_2}$$

9-7

# Mixed Stoichiometry Problems Answer Key

**Yan Bai**



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## **Table of Contents Mixed Stoichiometry Problems Answer Key**

1. Understanding the eBook Mixed Stoichiometry Problems Answer Key
  - The Rise of Digital Reading Mixed Stoichiometry Problems Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Mixed Stoichiometry Problems Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mixed Stoichiometry Problems Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Mixed Stoichiometry Problems Answer Key
  - Personalized Recommendations
  - Mixed Stoichiometry Problems Answer Key User Reviews and Ratings
  - Mixed Stoichiometry Problems Answer Key and Bestseller Lists
5. Accessing Mixed Stoichiometry Problems Answer Key Free and Paid eBooks

- Mixed Stoichiometry Problems Answer Key Public Domain eBooks
- Mixed Stoichiometry Problems Answer Key eBook Subscription Services
- Mixed Stoichiometry Problems Answer Key Budget-Friendly Options
- 6. Navigating Mixed Stoichiometry Problems Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Mixed Stoichiometry Problems Answer Key Compatibility with Devices
  - Mixed Stoichiometry Problems Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Mixed Stoichiometry Problems Answer Key
  - Highlighting and Note-Taking Mixed Stoichiometry Problems Answer Key
  - Interactive Elements Mixed Stoichiometry Problems Answer Key
- 8. Staying Engaged with Mixed Stoichiometry Problems Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Mixed Stoichiometry Problems Answer Key
- 9. Balancing eBooks and Physical Books Mixed Stoichiometry Problems Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Mixed Stoichiometry Problems Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Mixed Stoichiometry Problems Answer Key
  - Setting Reading Goals Mixed Stoichiometry Problems Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mixed Stoichiometry Problems Answer Key
  - Fact-Checking eBook Content of Mixed Stoichiometry Problems Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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