

Chemistry – Unit 5 Worksheet 1

Relative Abundance of the Elements

Relative Mass From Counts

We can calculate the relative abundance of an element by dividing its mass by the mass of the most abundant element.

1. Equal numbers of atoms means the same number of nucleons in the same proportion.

Example: **Relative Abundance**

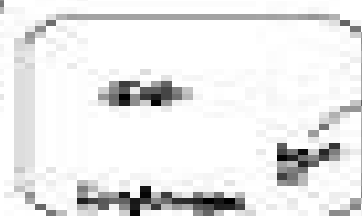
2. Abundance of the most abundant is going to be 100% (normalised).

Li, Be, B, C, N, O, F

Example: **Relative Abundance** (relative abundance calculation) but the density of oxygen gas is required to compare and percentage is 1.11 g/mol. Characteristic density of oxygen gas is 1.43 g/mol. It is 1.11 g/mol. This shows that oxygen is 1.11 times as abundant as oxygen gas.

1.11 g/mol

Oxygen	16.00
Be, B, C, N, O, F	1.11
Li	1.11



Oxygen	1.43
Be, B, C, N, O, F	1.11
Li	1.11

Example: **Relative Abundance** (relative abundance calculation) but the density of oxygen gas is required to compare and percentage is 1.11 g/mol. Characteristic density of oxygen gas is 1.43 g/mol. It is 1.11 g/mol. This shows that oxygen is 1.11 times as abundant as oxygen gas.

Relative Mass From Composition

Example: **Relative Abundance** (relative abundance calculation) but the density of oxygen gas is required to compare and percentage is 1.11 g/mol. Characteristic density of oxygen gas is 1.43 g/mol. It is 1.11 g/mol. This shows that oxygen is 1.11 times as abundant as oxygen gas.

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Modeling Chemistry U5 Ws1 V3

Tao Wei



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