

## Chemistry – Unit 5 Worksheet 1

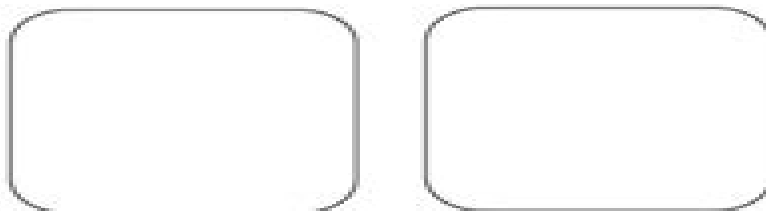
### Molar Masses of the Elements

#### Relative Mass From Gases

We have established that the combining ratio of gases can be explained if two assumptions are made:

1. Equal volumes of gases contain the same number of molecules at the same pressure and temperature.
2. Some pure elemental gases are clustered into pairs to form diatomic molecules.

Use these assumptions and particle diagrams to explain the fact that the density of oxygen gas at standard temperature and pressure is 1.43 g/liter, whereas the density of hydrogen gas under these conditions is 0.089 g/liter. How many times more massive is one molecule of oxygen than one molecule of hydrogen?



You shouldn't conclude that chemists were able to determine the molar masses of all the elements using this technique. Measurements of the density of the gaseous phase of many of the elements are difficult, if not impossible. However, we are going to see that chemists could use another tool – the *percent composition of compounds* to determine molar masses.

#### Relative Mass From Compounds

Many substances combine with oxygen to form a type of compound called an oxide. In the previous unit we saw that such combinations often occur in multiple proportions. John Dalton made the assumption that the lowest ratio was a 1:1 combination of elements. For now, we will make a similar assumption. *We may have to re-examine this assumption later.*

In Unit 4 you examined % composition data of several compounds. For example, 100 g of the simplest oxide of carbon contains 42.9g of C and 57.1 g of O. To be able to compare masses of elements to the mass of oxygen, it is important to use the same mass of oxygen in all of the ratios. Using a proportion, one can show that 75.0 g of C combines with 100 g of oxygen.

$$\frac{42.9\text{gC}}{57.1\text{gO}} = \frac{75.0\text{gC}}{100\text{gO}}$$

# Modeling Chemistry Unit 1

**Ensheng Dong**



## **Modeling Chemistry Unit 1:**

The book delves into Modeling Chemistry Unit 1. Modeling Chemistry Unit 1 is a crucial topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Modeling Chemistry Unit 1, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Modeling Chemistry Unit 1
- Chapter 2: Essential Elements of Modeling Chemistry Unit 1
- Chapter 3: Modeling Chemistry Unit 1 in Everyday Life
- Chapter 4: Modeling Chemistry Unit 1 in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Modeling Chemistry Unit 1. The first chapter will explore what Modeling Chemistry Unit 1 is, why Modeling Chemistry Unit 1 is vital, and how to effectively learn about Modeling Chemistry Unit 1.
3. In chapter 2, the author will delve into the foundational concepts of Modeling Chemistry Unit 1. The second chapter will elucidate the essential principles that must be understood to grasp Modeling Chemistry Unit 1 in its entirety.
4. In chapter 3, this book will examine the practical applications of Modeling Chemistry Unit 1 in daily life. The third chapter will showcase real-world examples of how Modeling Chemistry Unit 1 can be effectively utilized in everyday scenarios.
5. In chapter 4, the author will scrutinize the relevance of Modeling Chemistry Unit 1 in specific contexts. The fourth chapter will explore how Modeling Chemistry Unit 1 is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, this book will draw a conclusion about Modeling Chemistry Unit 1. The final chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Modeling Chemistry Unit 1.

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