STATION 3 . POLARITY

PH H S	CF. F. C. F.	H:8:H
Y0 10.7	MON-120 LAY	non-poles
: Q:: \$: 0:		: 3): 3: CI-
Polar	hoh-polar	Polor

7 · Molecular Structure

STATION 4 . IMF's AND ATTRACTIONS

Match	the st	atement with the Inter-l	Pourt	icle Forces		
		le-Dipole Interactions		Covalent bonds	- 0	. Hydrogen-bonding
D.	. Jones	c Bonds	E.	Metallic Bonds	E	. London Dispersion Forces
F	. 1.	Is used to explain why	BF	of Xe > Kr > Ar > N	le > H	
18	2.	Is present in Complant a	ed:	Cotamost)		
S	3.	Is used to explain why	Cu	is a good conductor		
10	4.	Is used to explain why	No	Cl(l) is a good condu	ictor	
A	5.	Is used to explain why	IC	l has a higher BP ther	n Bea	
	6.	Is used to explain why	bo	iling point of CH ₅ OH	is big	her than CH ₆ .
F	7.	A lattice of positive io	os i	n a sea of delocalized	l electr	rons
B	8.	Positive ends of molec	ule	s attract negative end	is of or	her molecules
E	9.	Lattice of alternating p	eo sá	tively and negatively	churg	ed particles
C	10.	Hydrogen atoms tether	red	between small, highl	y elect	ronegative atoms N, O, and F.
5	11.	Polarized electron clos	adis	induce dipoles in the	ir neig	aboring atoms

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RS Peters

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