

Unit 3 Cellular Energetics AP Biology Exam Review

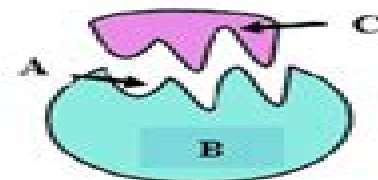
Enzyme Structure

1) To the right is an enzyme-substrate complex. Label what A, B, and C represent.

A: active site, B: enzyme, C: substrate

2) Describe how enzymes work. Make sure to include the terms products, reactants, and activation energy.

Enzymes are catalysts that speed up chemical reactions by binding to specific reactant molecules, or substrates, at their active sites. By binding to substrates, enzymes lower the activation energy required for the reaction to proceed, facilitating the conversion of reactants into products. Once the reaction is complete, enzymes release the products and can catalyze additional reactions.



Enzyme Catalysis and Environmental Effects

3) How can a scientist make an enzyme reaction increase in production rates?

A scientist can increase enzyme reaction rates by optimizing pH and temperature, increasing substrate or enzyme concentration, using cofactors, immobilizing enzymes, or genetically engineering them for improved efficiency.

4) What could slow down the creation of products from an enzymatic reaction? Give at least three items.

Factors that could slow down enzymatic reaction rates include suboptimal pH or temperature, insufficient substrate concentration, and enzyme inhibition.

5) What is denaturation? What causes this process to occur?

Denaturation is the process by which a protein, including enzymes, loses its natural shape and function due to external factors like heat, pH extremes, or chemicals. This disruption of the protein's structure prevents it from binding to its substrate and catalyzing reactions effectively.

pH Math

4) What's the concentration of a HCl solution that has a pH of 3.5?

$3.16 \times 10^{-4} \text{ M}$

Cell Energy: First and Second Law of Thermodynamics

5) Compare the first law to the second law of thermodynamics. Give examples of each.

The first law of thermodynamics, also known as the law of conservation of energy, states that energy cannot be created or destroyed, only transformed from one form to another. An example of this is the conversion of chemical energy in food into mechanical energy by muscles during exercise.

On the other hand, the second law of thermodynamics states that the entropy, or disorder, of an isolated system tends to increase over time. This law implies that spontaneous processes in a system will always lead to an increase in entropy. An example is the dissipation of heat from a hot object to its surroundings, leading to an overall increase in entropy.

$$\text{pH} = -\log[\text{H}^+]$$
$$[\text{H}^+] = 10^{-\text{pH}}$$

Modern Biology 3 Review Answers

ML Morrison



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