## 17.2 Measuring and Expressing Enthalpy Changes

# Enthalpy Change in a Calorimetry Experiment

When 25.0 mL of water containing 0.025 mol HCI at 25.0°C is added to 25.0 mL of water containing 0.025 mol NaOH at 25.0°C in a foam-cup calorimeter, a reaction occurs. Calculate the enthalpy change (in kJ) during this reaction if the highest temperature observed is 32.0°C. Assume that the densities of the solutions are 1.00 g/mL and the volume of the final solution is equal to the sum of the volumes of the reacting solutions.

### **Measuring And Expressing Enthalpy Changes Answers**

**Kathleen Armour** 

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ways detailed introduction to current knowledge techniques and applications The volume is essentially in two parts the first half covering a basic introduction to proteins appropriate at the undergraduate and early postgraduate level which will prove a valuable teaching aid The second half is a more advanced guide to concepts and methods covering a range of aspects not previously collected in one volume It will serve as a background reader and guide for advanced research study General Chemistry Raymond Chang, 2003 Publisher Description Automatic Potentiometric Titrations G. Svehla, 1978 International Series in Analytical Chemistry, 1975 General Chemistry Darrell D. Ebbing, Larry K. Krannich, Joan I. The Nature of Seawater Edward D. Senyk, 1996 Brotherhood of Locomotive Engineers' Monthly Journal, 1937 Goldberg, 1975 **Potassium in Soil** International Potash Institute, 1972 Random House Word Menu Stephen Glazier, 1997 A revolutionary tool that has changed the way we use words the Random House Word Menu functions in four ways it is a thesaurus with definitions a dictionary divided into word categories a reverse dictionary and a collection of glossaries A writer's right hand and a browser's delight this reference contains thousands of entries in over 800 categories **British Chemical Engineering**, 1969 British Chemical Abstracts, 1943 **British Abstracts** .1942 **British** Chemical and Physiological Abstracts ,1943 Starch and Its Derivatives Jack Augustus Radley, 1968

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