

Experiment 5

Determination of the Percentage of Ligands in $K_3[Fe(C_2O_4)_3] \cdot 3H_2O$

OBJECTIVES

1. To standardise the potassium permanganate solution.
2. To determine the percentage of ligands in $K_3[Fe(C_2O_4)_3] \cdot 3H_2O$.

INTRODUCTION

Standardisation of potassium permanganate ($KMnO_4$) solution is carried out by titration against a standard solution of oxalic acid. Oxalic acid is a good primary standard because the compound is available in solid form, as $H_2C_2O_4 \cdot 2H_2O$, which can be prepared to a very high degree of purity and is not hygroscopic nor efflorescent.

The reaction is as follows:



OVERALL REACTION



The following points about this reaction should be noted:

1. The reaction occurs only in acidic solutions. Acidification is achieved by the addition of sulfuric acid.
2. The reaction rate is normally slow. However, the reaction rate in this experiment is increased by (i) warming the contents of the conical flask or (ii) catalysed by the presence of manganese(II) ions, which is a product of the reaction.

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