



Ninja ZX-11
ZZ-R1100



Motorcycle Service Manual

Kawasaki Ninja Zx 11 Zz R1100 1993 2001 Service Manual

Kawasaki Jūkōgyō Kabushiki Kaisha



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water $n_a = 54 \text{ g} / 18 \text{ g mol}^{-1} = 3 \text{ mol}$ total number of moles $n_a + n_b = 0.5 + 3 = 3.5 \text{ mol}$ mole fraction of solute ethyl alcohol $x_b = n_b / n_a + n_b = 0.5 / 3.5 = 0.1429$

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40 g of H_2O molar mass of H_2O is 18 g/mol and molar mass of CH_3OH is 32 g/mol solution moles of CH_3OH $45 / 32 = 1.4$
moles of H_2O $40 / 18 = 2.2$ moles therefore according to the equation

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time the total pressure $p_a = x_a p_t$ mole fraction $x_a = \frac{\text{moles } a}{\text{moles total}}$ $\sum_{i=1}^n x_i = 1$ mole fraction of substance

x_a is often denoted by x_a and

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by 100 gives the mole percentage 3 mole fraction is a unitless and dimensionless expression advantages of molar fraction

mole fraction is not dependent on the temperature

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moles times molar mass equals grams 1 31108 mol 119 0023 g mol 156 g kbr

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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xylene c 8 h 10 has a density of 0.864 g/ml a solution was prepared by mixing acetone and xylene calculate the mole fraction of acetone in a solution if 19.0 ml of acetone was dissolved in 98.4 ml of xylene 7.1

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web since the mole fraction of the O_2 is 0.050 it contributes 0.050 or if you will 5.0 of the total pressure thus total pressure is 0.21 0.050 4.2 atm problem 9 a sample of 1.43 g of helium and an unweighed quantity of O_2 are mixed in a flask at room temperature