

5 Descriptions of the bonding in three substances are given.

- substance 1 strong covalent bonds between atoms, permanent dipole-dipole attractions between molecules
- substance 2 strong covalent bonds between atoms, weak forces between molecules
- substance 3 strong covalent bonds between atoms, hydrogen bonding between molecules

Which compounds could be substances 1, 2 and 3?

|          | substance 1                     | substance 2                     | substance 3                     |
|----------|---------------------------------|---------------------------------|---------------------------------|
| <b>A</b> | CH <sub>3</sub> OH              | Al <sub>2</sub> Cl <sub>6</sub> | CH <sub>2</sub> Cl <sub>2</sub> |
| <b>B</b> | Al <sub>2</sub> Cl <sub>6</sub> | CH <sub>2</sub> Cl <sub>2</sub> | CH <sub>4</sub>                 |
| <b>C</b> | CH <sub>2</sub> Cl <sub>2</sub> | CH <sub>4</sub>                 | CH <sub>3</sub> OH              |
| <b>D</b> | CH <sub>4</sub>                 | CH <sub>3</sub> OH              | H <sub>2</sub> O                |

6 Which type of bonding is **never** found in elements?

- A** covalent
- B** ionic
- C** metallic
- D** van der Waals' forces

7 Using the information in the table, what is the enthalpy change,  $q$ , for the reaction described?



| standard enthalpy change                           | value / kJ mol <sup>-1</sup> |
|----------------------------------------------------|------------------------------|
| $\Delta H_{\text{f}}^\circ \text{ CsBr}(\text{s})$ | +25.9                        |
| $\Delta H_{\text{f}}^\circ \text{ Cs}^+(\text{g})$ | -276                         |
| $\Delta H_{\text{f}}^\circ \text{ Br}^-(\text{g})$ | -335                         |

- A** -636.9      **B** -585.1      **C** +585.1      **D** +636.9

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