

|  |  |  |
|--|--|--|
| <b>Activity A:</b><br><b>Small atoms</b> | <b>Get the Gizmo ready:</b> <ul style="list-style-type: none"> <li>On the PERIODIC TABLE tab, select H (hydrogen).</li> <li>Select the ELECTRON CONFIGURATION tab.</li> <li>Click <b>Reset</b>.</li> </ul> |  |
|--|--|--|

**Introduction:** Electrons are arranged in **orbitals**, **subshells**, and **shells**. These levels of organization are shown by the boxes of the Gizmo. Each box represents an orbital. The subshells are labeled with letters (*s*, *p*, *d*, and *f*) and the shells are labeled with numbers.

**Question:** How are electrons arranged in elements with atomic numbers 1 through 10?

1. Infer: Based on its atomic number, how many electrons does a hydrogen atom have? \_\_\_\_\_

2. Arrange: The **Aufbau principle** states that electrons occupy the lowest-energy orbital. Click once in the **1s** box to add an electron to the only orbital in the *s* subshell of the first shell.

Click **Check**. What is the electron configuration of hydrogen? \_\_\_\_\_

3. Arrange: Click **Next element** to select helium. Add another electron to the **1s** orbital. The arrows represent the **spin** of the electron. What do you notice about the arrows?

\_\_\_\_\_

The **Pauli exclusion principle** states that electrons sharing an orbital have opposite spins.

4. Check your work: Click **Check**. What is the electron configuration of helium? \_\_\_\_\_

5. Arrange: Click **Next element** and create electron configurations for lithium, beryllium, and boron. Click **Check** to check your work, and then list each configuration below:

Lithium: \_\_\_\_\_ Beryllium: \_\_\_\_\_ Boron: \_\_\_\_\_

6. Arrange: Click **Next element** to select carbon. Add a second electron to the first **2p** orbital.

Click **Check**. What feedback is given? \_\_\_\_\_

7. Rearrange: **Hund's rule** states that electrons will occupy an empty orbital when it is available in that subshell. Rearrange the electrons within the **2p** subshell and click **Check**.

Is the configuration correct now? \_\_\_\_\_

Show the correct configuration in the boxes at right:

1s ☐

2s ☐ 2p ☐ ☐ ☐

(Activity A continued on next page)

# Ionic Bonds Explore Learning Answer Key

**N Colangelo**



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