

Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
Li	1 ve	transfers $1e^-$ to F	Li^+
F	7 ve	accepts $1e^-$ from Li	F^-

Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) $Li + Cl \Rightarrow LiCl$	Li	1 ve	donates $1e^-$	Li^+
$Li + Cl \Rightarrow LiCl$	Cl	1 ve	accepts $1e^-$	Cl^-
2) $Ca + O \Rightarrow CaO$	Ca	2 ve	donates $2e^-$	Ca^{+2}
$Ca + O \Rightarrow CaO$	O	6 ve	accepts $2e^-$	O^{+2}
3) $Be + F \Rightarrow BeF_2$	Be	2 ve	donates $2e^-$ to each F	Be^{+2}
$Be + F \Rightarrow BeF_2$	F	7 ve	each F accepts $1e^-$ each	F^-
4) $Mg + S \Rightarrow MgS$	Mg	2 ve	donates $2e^-$ to S	Mg^{+2}
$Mg + S \Rightarrow MgS$	S	6 ve	accepts $2e^-$	S^{+2}
5) $K + F \Rightarrow KF$	K	1 ve	donates $1e^-$	K^+
$K + F \Rightarrow KF$	F	7 ve	accepts $1e^-$	F^-

Ionic Bonding Advanced Chemquest Answers

LL Leslie



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Table of Contents Ionic Bonding Advanced Chemquest Answers

1. Understanding the eBook Ionic Bonding Advanced Chemquest Answers
 - The Rise of Digital Reading Ionic Bonding Advanced Chemquest Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Bonding Advanced Chemquest Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ionic Bonding Advanced Chemquest Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic Bonding Advanced Chemquest Answers
 - Personalized Recommendations
 - Ionic Bonding Advanced Chemquest Answers User Reviews and Ratings
 - Ionic Bonding Advanced Chemquest Answers and Bestseller Lists
5. Accessing Ionic Bonding Advanced Chemquest Answers Free and Paid eBooks

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 - Ionic Bonding Advanced Chemquest Answers eBook Subscription Services
 - Ionic Bonding Advanced Chemquest Answers Budget-Friendly Options
6. Navigating Ionic Bonding Advanced Chemquest Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Ionic Bonding Advanced Chemquest Answers Compatibility with Devices
 - Ionic Bonding Advanced Chemquest Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ionic Bonding Advanced Chemquest Answers
 - Highlighting and Note-Taking Ionic Bonding Advanced Chemquest Answers
 - Interactive Elements Ionic Bonding Advanced Chemquest Answers
 8. Staying Engaged with Ionic Bonding Advanced Chemquest Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ionic Bonding Advanced Chemquest Answers
 9. Balancing eBooks and Physical Books Ionic Bonding Advanced Chemquest Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ionic Bonding Advanced Chemquest Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Ionic Bonding Advanced Chemquest Answers
 - Setting Reading Goals Ionic Bonding Advanced Chemquest Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Ionic Bonding Advanced Chemquest Answers
 - Fact-Checking eBook Content of Ionic Bonding Advanced Chemquest Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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