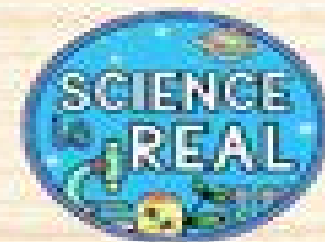


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TYPES OF CHEMICAL BONDS
Atomic Combinations

1. IONIC BONDING: Occurs between a metal and a non-metal. Electrons are transferred from the metal to the non-metal, forming oppositely charged ions. These ions are held together by strong electrostatic forces of attraction, forming an ionic bond. Example: Sodium Chloride (NaCl).

2. COVALENT BONDING: Occurs between two non-metals. Atoms share their outer shell electrons to form a strong bond. Example: Water (H₂O).

3. METALLIC BONDING: Occurs between metal atoms. The atoms are held together by a sea of delocalized electrons. Example: Iron (Fe).

ATOMIC STRUCTURE

Atoms are made up of three subatomic particles: protons, neutrons, and electrons. Protons and neutrons are located in the nucleus, while electrons orbit the nucleus in shells. The number of protons in an atom is equal to the number of electrons, making the atom electrically neutral.

PERIODIC TABLE

The periodic table is a chart that organizes elements based on their atomic number and chemical properties. It is divided into groups and periods. Elements in the same group have similar properties.

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