

Chemistry: Ionic & Covalent Bonds Unit Study Guide

Covalent Bonding & Nomenclature Study Guide

1. Circle the correct word that would complete the statements below regarding Covalent Compounds.

- Bondable
- Large / Small
- Always / Never
- High / Low
- Conducts / Does not Conduct Electricity
- Solid / Liquid / Gas at Room Temperature

2. Write the prefix following number.

Number	Prefix
1	
2	
3	
4	
5	

3. What are the two

4. Complete the chart

Formula	Name
SO_2	
	Carbon
H_2O	

5. What does VSEPR

Name: _____ Date: _____ Period: _____

Ionic Bonding & Nomenclature Study Guide

1. Circle the correct word that would complete the statements below regarding Ionic Compounds.

- Bond between a Metal / Nonmetal and a Metal / Nonmetal
- Large / Small Electronegativity Difference
- The (Metal / Nonmetal) form Cations which (Gain / Lose) Electrons
- The (Metal / Nonmetal) form Anions which (Gain / Lose) Electrons
- High / Low Melting & Boiling Point
- Conducts / Does not Conduct Electricity in Solution
- Solid / Liquid / Gas at Room Temperature

2. What charge would the following periodic table families have in a chemical bond?

Family	Charge
Alkali Metals	
Alkaline Earths	
Transition Metals	
Carbon	
Nitrogen	
Halogens	

3. Write the roman numeral that is represented by the following numbers.

Roman Numeral	Number
	1
	2
	3
	4
	5
	6

4. Complete the chart below by filling in the missing names.

Cation	Anion	Formula	Name
Mg^{2+}	Br^-		
		CuO	
			Iron (II) chloride
		$(NH_4)_3PO_4$	

Ionic Covalent Bonding Test Study Guide Answers

Susan McMurry



Ionic Covalent Bonding Test Study Guide Answers:

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questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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