



# CHEMICAL BONDS



## IONIC, COVALENT, & METALLIC

**FORMS OF CHEMICAL BONDS**  
Bonding: combinations

1. **IONIC BONDING** occurs between a metal and a nonmetal. The metal atom loses one or more electrons to become a positively charged cation, and the nonmetal atom gains those electrons to become a negatively charged anion. The resulting electrostatic attraction between the oppositely charged ions forms the ionic bond.

2. **COVALENT BONDING** occurs between two nonmetals. The atoms share one or more pairs of electrons to achieve a full outer shell. The shared pair of electrons is called a covalent bond.

3. **METALLIC BONDING** occurs between metal atoms. The metal atoms release their valence electrons into a 'sea of delocalized electrons' that are shared by all the metal ions in the lattice.

**SCIENCE REAL**

A simplified periodic table of elements, color-coded by groups.

**IONIC BONDING**

Diagram showing the formation of an ionic bond between a sodium atom (Na) and a chlorine atom (Cl). Sodium loses an electron to become Na<sup>+</sup>, and chlorine gains an electron to become Cl<sup>-</sup>. The resulting NaCl crystal lattice is shown as a 3D structure of alternating positive and negative ions.

**COVALENT BONDING**

Diagram showing the formation of a covalent bond between two hydrogen atoms (H). Each hydrogen atom has one electron. They share their electrons to form a pair, represented by a line between the two atoms, creating a stable H<sub>2</sub> molecule.

**METALLIC BONDING**

Diagram showing a metal lattice where metal ions are arranged in a regular pattern, and a 'sea of delocalized electrons' surrounds them. This model explains the properties of metals, such as conductivity and malleability.

## GUIDED READING ACTIVITY

# Ionic Bonds Chapter 5 Reading Notetaking Guide Answers

**Sabine Zange**



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1  $v_L$  t  $v_{g1}$  t  $r_L$  and when it is in position 2  $v_L$  t  $v_{t1}$  t  $r_L$  the inductor voltage  $v_L$  t is shown in figure 1 applying inductor volt  
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