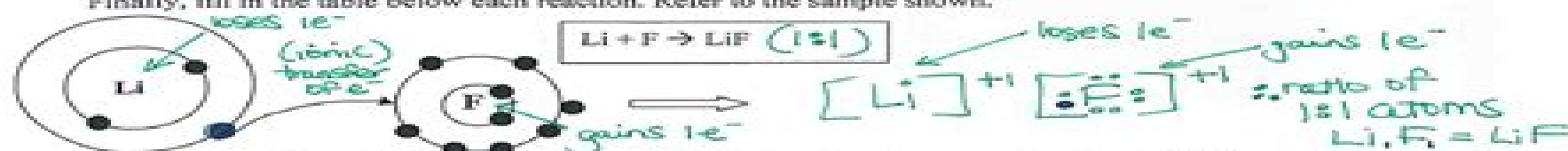


Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
Li	1 ve	transfers $1e^-$ to F	$Li^+$
F	7 ve	accepts $1e^-$ from Li	$F^-$

Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) $Li + Cl \Rightarrow LiCl$ $Li + \cdot\ddot{Cl}: \rightarrow [Li]^+ [\ddot{Cl}:]^{-1}$	Li	1 ve	donate $1e^-$	$Li^+$
	Cl	1 ve	accepts $1e^-$	$Cl^{-1}$
2) $Ca + O \Rightarrow CaO$ $Ca + \cdot\ddot{O}: \rightarrow [Ca]^{+2} [\ddot{O}:]^{-2}$	Ca	2 ve	donates $2e^-$	$Ca^{+2}$
	O	6 ve	accepts $2e^-$	$O^{-2}$
3) $Be + F \Rightarrow BeF_2$ $Be + 2 \cdot\ddot{F}: \rightarrow [Be]^{+2} [:\ddot{F}:]^{-1}_2$	Be	2 ve	donates $2e^-$ to each F	$Be^{+2}$
	F	7 ve	each F accepts $1e^-$ each	$F^{-1}$
4) $Mg + S \Rightarrow MgS$ $Mg + \cdot\ddot{S}: \rightarrow [Mg]^{+2} [\ddot{S}:]^{-2}$	Mg	2 ve	donates $2e^-$ to S	$Mg^{+2}$
	S	6 ve	accepts $2e^-$	$S^{-2}$
5) $K + F \Rightarrow KF$ $K + \cdot\ddot{F}: \rightarrow [K]^+ [:\ddot{F}:]^{-1}$	K	1 ve	donates $1e^-$	$K^{+1}$
	F	7 ve	accepts $1e^-$	$F^{-1}$

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**Clemens Wendtner**



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bondlines adhesion between compounded elastomers applications of contact angle measurements in pharmaceuticals and foods oxygen or ammonia plasma treatment of polyolefin surfaces surface free energy determination of powders and particles wood bonds and dispersion adhesion forces between macroscopic objects

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## **Table of Contents Ionic Bonds Reveiw And Reinforce Answers**

1. Understanding the eBook Ionic Bonds Reveiw And Reinforce Answers
  - The Rise of Digital Reading Ionic Bonds Reveiw And Reinforce Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Bonds Reveiw And Reinforce Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ionic Bonds Reveiw And Reinforce Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic Bonds Reveiw And Reinforce Answers
  - Personalized Recommendations
  - Ionic Bonds Reveiw And Reinforce Answers User Reviews and Ratings
  - Ionic Bonds Reveiw And Reinforce Answers and Bestseller Lists
5. Accessing Ionic Bonds Reveiw And Reinforce Answers Free and Paid eBooks

- Ionic Bonds Reveiw And Reinforce Answers Public Domain eBooks
  - Ionic Bonds Reveiw And Reinforce Answers eBook Subscription Services
  - Ionic Bonds Reveiw And Reinforce Answers Budget-Friendly Options
6. Navigating Ionic Bonds Reveiw And Reinforce Answers eBook Formats
    - ePub, PDF, MOBI, and More
    - Ionic Bonds Reveiw And Reinforce Answers Compatibility with Devices
    - Ionic Bonds Reveiw And Reinforce Answers Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Ionic Bonds Reveiw And Reinforce Answers
    - Highlighting and Note-Taking Ionic Bonds Reveiw And Reinforce Answers
    - Interactive Elements Ionic Bonds Reveiw And Reinforce Answers
  8. Staying Engaged with Ionic Bonds Reveiw And Reinforce Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Ionic Bonds Reveiw And Reinforce Answers
  9. Balancing eBooks and Physical Books Ionic Bonds Reveiw And Reinforce Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Ionic Bonds Reveiw And Reinforce Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Ionic Bonds Reveiw And Reinforce Answers
    - Setting Reading Goals Ionic Bonds Reveiw And Reinforce Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Ionic Bonds Reveiw And Reinforce Answers
    - Fact-Checking eBook Content of Ionic Bonds Reveiw And Reinforce Answers
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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