

How to Name Chemical Compounds

is it ionic? is an anion,
or does it contain a polyatomic ion (as provided by H⁺)?

Ionic compound

[cation's name] [anion's name]

main- element	<ul style="list-style-type: none"> take its element name! ex. Al^{3+}: aluminum if transition metal*, include charge in Roman numeral ex. Cr^{2+}: chromium (II) 	<ul style="list-style-type: none"> get element's name, chop end, add "-ide" ex. F^{-}: fluorine → fluoride ✓
poly- element	<ul style="list-style-type: none"> just use its name! ex. NH_4^{+}: ammonium 	<ul style="list-style-type: none"> just use its name! ex. CO_3^{2-}: carbonate

* exceptions (no Roman numerals):

silver ion: always Ag^{+}

zinc ion: always Zn^{2+}

cadmium ion: always Cd^{2+}

Molecular compound

is it HCO_3^{-} ? otherwise →

Acid! [1st element name] [2nd element root] + "-ide"

↑ skip if "mono-" ↑ skip if 1st element is H

Numerical prefixes

- | | |
|-------------|-------------|
| • 1: mono- | • 6: hexa- |
| • 2: di- | • 7: hepta- |
| • 3: tri- | • 8: octa- |
| • 4: tetra- | • 9: nona- |
| • 5: penta- | • 10: deca- |

for
first element: never
"mono-"
✓ principle

Acids

Ignore all H^{+} , and look at the name of the remaining anion.

Anion	X-ide	X-ate	X-ite
Acid	hydro-X-ic acid	X-ic acid	X-ous acid
	ex. Br^{-} bromide HBr hydrobromic acid	ex. $\text{C}_2\text{H}_3\text{O}_2^{-}$ acetate $\text{HC}_2\text{H}_3\text{O}_2$ acetic acid	ex. IO_3^{-} iodate HIO_3 iodic acid
	ex. CN^{-} cyanide HCN hydrocyanic acid	ex. PO_4^{3-} phosphate H_3PO_4 phosphoric acid	ex. SO_3^{2-} sulfite H_2SO_3 sulfurous acid

Ions And Ionic Compounds Study Guide

Peipei Pang



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