# Writing K<sub>a</sub>

 $HC_2H_3O_2(aq) + H_2O(I) \rightleftharpoons H_3O^+(aq) + C_2H_3O_2^-(aq)$ 

Which of the following is the correct K<sub>a</sub> expression for the reaction shown above?

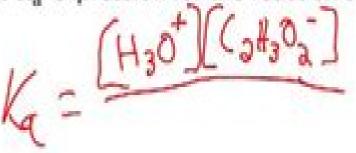
(A) 
$$K_a = \frac{[H_3O^+][C_2H_3O_2^-]}{[HC_2H_3O_2][H_2O]}$$

(B) 
$$K_a = \frac{[H_3O^+][C_2H_3O_2^-]}{[HC_2H_3O_2]}$$

(C) 
$$K_a = \frac{[C_2H_3O_2^-]}{[HC_2H_3O_2]}$$

(D) 
$$K_a = \frac{[H_3O^+]}{[H_2O]}$$

(E) 
$$K_a = \frac{[HC_2H_3O_2]}{[H_3O^+][C_2H_3O_2^-]}$$



# **Ka Practice Chemistry**

**M** Walker

**Ka Practice Chemistry:** 

This book delves into Ka Practice Chemistry. Ka Practice Chemistry is a crucial topic that must be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Ka Practice Chemistry, encompassing both the fundamentals and more intricate discussions.

- 1. This book is structured into several chapters, namely:
  - Chapter 1: Introduction to Ka Practice Chemistry
  - Chapter 2: Essential Elements of Ka Practice Chemistry
  - o Chapter 3: Ka Practice Chemistry in Everyday Life
  - Chapter 4: Ka Practice Chemistry in Specific Contexts
  - ∘ Chapter 5: Conclusion
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- 3. In chapter 2, the author will delve into the foundational concepts of Ka Practice Chemistry. This chapter will elucidate the essential principles that need to be understood to grasp Ka Practice Chemistry in its entirety.
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